

10. 11. 2019



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- ### 3. 研究設計

☞ কোলাস পানিসূত্র লবণের দ্রাব্যতা আপনাত্মা কৃষ্ণির সাথে সাথে এক পর্যায়ে কৃষ্ণি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা কৃষ্ণির সাথে সাথে দ্রাব্যতা কৃষ্ণি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কোলাস পানি দ্রুত লবণ নিষ্কৃতিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলার যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত হওয়ার সময় বিয়োজন প্রক্রিয়াটি তাপশরী বিক্রিয়া কিন্তু 32.4°C আপনাত্মার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C আপনাত্মার উপরে আপনাত্মা কৃষ্ণিজনিত কারণে দ্রাব্যতা হ্রাস ঘটে।

$w = \text{SMV}$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

দ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 দ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

$$\text{[3]} \quad \text{H}_2\text{SO}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{SO}_4 = \left(0.5 \times \frac{2.50}{1000}\right) \text{ mol} \\ = 0.125 \text{ mol}$$
$$\begin{aligned} \text{NaOH-H: Na}_2\text{CO}_3 \text{ से अभिक्रिया करने पर} &= \left(0.11 \times \frac{200}{1000} \right) \text{ mol} \\ &= 0.022 \text{ mol} \end{aligned}$$

অম্লটিত বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
 সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 0.077 mol " " " " (0.077) mol " "
 " " " " " 0.077 mol "

অনশিট H_2SO_4 এর পরিমাণ = $(0.125 - 0.077) \text{ mol}$
 = 0.048 mol

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে ০.৫
mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

উদাহরণ ২ একজন ছাত্র ল্যাবরেটরিতে X-সম্পত্তিতে 10 mg এর ওজনের পরিমাণ এবং Y-সম্পত্তিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর দ্রবণ তৈরি করতে চায়।

Abstract

- নোড কাকে বলে?
- Cl_2 অপেক্ষার— ব্যাখ্যা করো।
- উদ্ভীপকের এসিডটির মৌল্যর ঘনমাত্রা নির্ণয় করো।
- উদ্ভীপকের কোন পদার্থটি পরিবেশবান্ধব? বিশ্লেষণ করো।

२. नए व्यंजन का डिखना

উদাহরণ ১: নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

১৩। সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অণুটির হয়। যেহেতু Cl_1 এর তড়িৎ ঋণাত্মকতা 3.0 ও C_2 অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_1 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লিখিত মৌলিক অণুর ক্ষেত্রে এছাড়াও থাকে।

গ) উদ্ভীপকে বর্ণিত এসিডটি হলো 2% $\left(\frac{W}{V}\right)$ HCl।

2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—

100 mL. 5% HCl ଏବଂ ତରଳ 2g

$$\therefore 1000 \text{ mL} = \left(\frac{2 \times 1000}{100} \right) \text{ g} = 20 \text{ g}$$

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JG Myers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the JG Myers text.

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Cambridge Checkpoints HSC Chemistry 2017-19 Maureen Slade, Roger Slade, 2016-06-29 Excel HSC Chemistry C. M. Roebuck, 2003 *Excel HSC Chemistry* Jim Stamell, 2011 ISBN 9781741252996 AUTHOR Jim Stamell RRP 39 95 PAGES 428 pp SPECIFICATION Softcover perfect bound 280 mm x 210 mm STATUS New edition PUBLICATION DATE April 2008 The EXCEL HSC Chemistry guide is directly linked to the syllabus with every single dot point of the HSC Chemistry syllabus appearing in the margin of the book You can write in the guide so your study is focused and your notes are structured This guide comes in a brand new format that makes even better use of your study time up to date coverage of the core topics plus 3 Option topics Industrial Chemistry Shipwrecks Corrosion and Conservation and Forensic Chemistry this guide is organised just like the HSC syllabus so the students learn to section the theoretical part is under routine headings and the students section the practical part is under headings like First hand Second hand and Investigations and Problem Solving %this way you will be able to see at a glance what the theoretical and practical work is all main headings in each chapter 1 1 2 1 etc are directly from the syllabus word for word %this way you can easily match the Excel guide to the syllabus an alphabetical list of all the key definitions and concepts you should know from each chapter %an efficient way of learning all the definitions in one go chapter syllabus checklist with every single dot point listed in checklist form for each chapter %a fantastic way of testing that you know all the work hundreds of key concept questions with answers %questions that test your recall of knowledge in each chapter HSC type questions for every section in each chapter with clock icons to tell you how much time you will have to answer the questions in the HSC %this way you can test yourself on HSC type questions under HSC type time pressure an examiner maximiser feature ticks to show the mark distribution and answers to all HSC type questions %all you need to answer HSC type questions two sample HSC papers with an examiner maximiser feature plus answers %not one but two up to date sample papers the Excel syllabus summary notes a detachable section at the end of the guide where every single dot point of each chapter is summarised for you% a comprehensive and compact summary of the whole course in 32 pages **Excel Revise in a Month HSC Chemistry** Will Marchment, 2007 *Excel Senior High School Hospitality* Cathy Smith, Georgina Buttel, Sonya Ankerholz-Potts, 2004 **Design and Technology** Deborah Trevallion, 2003 A Compendium of Geochemistry Yuan-Hui Li, 2000-10-22 A general understanding of these principles and processes including those pertaining to cosmology geology and biology is essential maintains the author for deciphering and predicting transport pathways and final sinks of anthropogenic pollutants in our environment BOOK JACKET

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