SOLUTIONS TO PROBLEMS

Chapter 1

- L. An amino or imino group gives a molecule a positive charge; a carbenylate or phosphoryl group gives a molecule a negative charge.
- A. Thiol isulthydrylls group.
 - B Carbonyl group
 - C. Armide linkage
 - D. Phosphounhydride (pyrophosphoryl) linkage
 - E. Phosphoryl group (P.):
 - F. Hlydroxyl group.
- 3. Hydrolysis reactions tend to occur with an increase in entropy, because the highly ordered polymer is broken down into separate units.
- Typical prokaryotic cells are 1–10 µm in diameter, so T. namibiousis is about 10-300 times larger (or has 1000 to 27 million times the volume). Typical eukaryotic cells are 10-100 µm in diameter, so-T. numblemais is about the same size or up to 30 times larger (or has the same to 27,000 times the volume).
- 5. The data indicate that eukaryotes are more similar to the archaea thus to the eubacteria.
- 6. (a) Liquid water, (b) ice has less entropy at the lower temperature.
- 7. (a) Decreases, (b) increases, (c) increases, (d) no change.
- 8. No. When the change in enthalpy is positive and the change in entropy is negative, the free energy change for the process is greater than zero, which makes the process unfavorable.
- 9. (a) False. A spontaneous reaction occurs in only one direction.
 - (b) False. Thermodynamics does not specify the rate of a reaction.
 - (c) True. (d) True. A reaction is spontaneous as long as ΔS > ΔWT.
- 10. Gas T = 273 + 10 = 283 K.
 - $\Delta G = \Delta H T\Delta S$ $\Delta G = 15 \text{ kJ} + (283 \text{ K})(0.05 \text{ kJ} \cdot \text{K}^{-1})$
 - $-15 14.15 \, kJ 0.85 \, kJ$

 ΔG is greater than zero, so the reaction is not spontaneous.

- (b) T 273 + 80 353 K
- $\Delta G = \Delta M T\Delta S$
- $\Delta G = 15 \text{ kJ} (353 \text{ K}) \cdot 00.05 \text{ kJ} \cdot \text{K}^{-1}$
 - -15 17.65 kJ -2.65 kJ

ΔεG is less than zero, so the reaction is spontaneous.

- II. $\Delta G = \Delta H T\Delta S$
 - $\Delta G = -7000 \text{ J} \cdot \text{mol}^{-1} (298 \text{ K})(-25 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1})$
 - $\Delta G = -7000 + 7450 J \cdot mol^{-1} = 450 J \cdot mol^{-1}$

The reaction is not spontaneous because $\Delta G > 0$. The temperature must be decreased in order to decrease the value of the $T\Delta S$ term.

12.
$$\Delta G^{**} = -RT \ln \frac{\{C\}}{1\Delta L(B)}$$

- $= -18,000 \text{ J} \cdot \text{mol}^{-1} = -18 \text{ kJ} \cdot \text{mol}^{-1}$
- 13. $\Delta G^{in} = -RT \ln K_{in} = -RT \ln(|C||D|/|A||B|)$
 - -(8.314.3 · K ⁻¹ · mol ⁻¹) (298 K) ln((3 × 10 ⁻¹) (5 × 10 ⁻¹) (10 × 10 °) (15 × 10 °)
 - 5700 J mol * 5.7 kJ mol *

Since ΔG^{**} is positive, the reaction is endergonic under standard conditions.

- 14. $K_{\infty} = e^{-2G^*/RT} = e^{-(-200000 \cdot \text{mod}^*)(0.7041 \cdot \text{K}^{-1} \cdot \text{mod}^*)(200 \text{K})}$
- -4.6×10^{5}
- From Eq. 1-17, K_{eq} = [G6P]/[G1P] = e⁻³⁶⁷⁻¹⁶⁷

[G6P]/[G1P] = e^-c-Trible and "planters - K" - mel" scores KG6PMDGHP1 - 17.6

 $1GIPPDG6P1 \rightarrow 0.057$

- 16. The cell membrane must be semipermeable so that the cell can retain executal compounds while allowing nutrients to enter and
- 17. Membrane-enclosed cellular compartments are typical of enkaryotic but not prokaryotic cells.
- Concentration (number of moles)/(volume)

Volume = $(4/3) \propto r^3 = (4/3) \times (5 \times 10^{-7} \text{ m})^3$ $-5.24 \times 10^{-10} \,\text{m}^3 - 5.24 \times 10^{-10} \,\text{L}$

Moles of protein = (2 molecules)/(6.022 × 10²³ molecules · mol⁻¹) - 3.32 × 10⁻²⁴ mod

Concentration = $(3.32 \times 10^{-14} \text{ moly}(5.24 \times 10^{-16} \text{ L})$ $-6.3 \times 10^{-9} M - 6.3 \text{ eM}$

 Number of molecules = (molar conc.)(volume) (6.022 × 10²³ molecules * mol⁻¹3

> $= (1.0 \times 10^{-3} \text{ mod} \cdot \text{L}^{-1})(5.24 \times 10^{-16} \text{L})$ (6.022 × 10¹³ molecules - mol⁻¹)

 -3.2×10^6 molecules

 In order for \(\Delta G\) to have a negative value (a spontaneous reaction). $T\Delta S$ must be greater than ΔH .

 $T\Delta S \ge \Delta M$

 $T \gg \Delta M \Delta S$

 $T > 7000 \text{ J} \cdot \text{mol}^{-1}/20 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1}$

F 15 350 K or 7770

21. (a) Because $K_{eq} = \frac{|\mathbb{R}|}{|\Omega|} = 25$, at equilibrium, the concentration of

R is 25 times greater than the concentration of Q. When equal concentrations of O and R are mixed, molecules of O will be converted to molecules of R.

(b) Let x = amount of Q converted to R, so that [R] will be 50 μM +

x and [Q] will be $50 \mu M - x$. Since $K_{eq} = \frac{\{R\}}{\{O\}^2} = 25$.

$$50 + x = 25(50 - x)$$

 $50 + x = 1250 - 25x$

 $26\pi - 1200$

x = 46.15 $[R] = 50\mu M + 46.15\mu M = 96.15\mu M$

 $||Q|| = 50 \mu M - 46.15 \mu M = 3.85 \mu M$

22. At 10°C = 283 K (1/T = 0.00353), K_{eq} = 100 and ln K = 4.61. At 30°C = 303 K (1/T = 0.00330), $K_{eq} = 10$ and In K = 2.30.

These two points generate a line on a van't Bloff plot (In K., versus L/T) with a positive slope that is equal to $(-\Delta M'/00)$. ΔM must therefore be negative, indicating that enthalpy decreases during the reaction (heat is given off).

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Michael Hoelscher

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