

Equilibrium

At what point is a reversible reaction “completed”?

Why?

Most of the reactions that we have studied this year have been “forward” reactions—once the reactant has changed into the product it stays that way. We can assume that eventually the reaction will “finish” when the limiting reactant runs out. However, in a reversible reaction (one that can take place in both directions), once the product is formed it can turn back into the reactant and a continuous cycle occurs. When do reactions such as these “finish”? Is there ever an end, and can we predict how much reactant or product will be present when the reaction is “done?”

Model 1 – A Reversible Reaction



1. What is the reactant in the reaction in Model 1?

A

2. What is the product in the reaction in Model 1?

B

3. What is the significance of the double arrow in the equation in Model 1?

The reaction is reversible, which means that changes can take place in both the forward and backward or reverse directions.



4. Imagine that the reaction in Model 1 starts with 100 molecules of A and zero molecules of B. Explain why the concentration of substance A will never reach zero.

As the molecules of A react to make B, some of the B molecules will react to make more A. So substance A is always being replenished.

5. Imagine the reaction in Model 1 starts with zero molecules of A and 100 molecules of B. Will the concentration of substance B increase or decrease as the reaction proceeds? Explain.

Since there are no A molecules, the concentration of B cannot increase. The reaction can only go in reverse initially. Therefore, the concentration of B will decrease until it reaches a stable value.

Equilibrium Pogil Answer

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POGIL Shawn R. Simonson, 2023-07-03 Process Oriented Guided Inquiry Learning POGIL is a pedagogy that is based on research on how people learn and has been shown to lead to better student outcomes in many contexts and in a variety of academic disciplines Beyond facilitating students mastery of a discipline it promotes vital educational outcomes such as communication skills and critical thinking Its active international community of practitioners provides accessible educational development and support for anyone developing related courses Having started as a process developed by a group of chemistry professors focused on helping their students better grasp the concepts of general chemistry The POGIL Project has grown into a dynamic organization of committed instructors who help each other transform classrooms and improve student success develop curricular materials to assist this process conduct research expanding what is known about learning and teaching and provide professional development and collegiality from elementary teachers to college professors As a pedagogy it has been shown to be effective in a variety of content areas and at different educational levels This is an introduction to the process and the community Every POGIL classroom is different and is a reflection of the uniqueness of the particular context the institution department physical space student body and instructor but follows a common structure in which students work cooperatively in self managed small groups of three or four The group work is focused on activities that are carefully designed and scaffolded to enable students to develop important concepts or to deepen and refine their understanding of those ideas or concepts for themselves based entirely on data provided in class not on prior reading of the textbook or other introduction to the topic The learning environment is structured to support the development of process skills such as teamwork effective communication information processing problem solving and critical thinking The instructor s role is to facilitate the development of student concepts and process skills not to simply deliver content to the students The first part of this book introduces the theoretical and philosophical foundations of POGIL pedagogy and summarizes the literature demonstrating its efficacy The second part of the book focusses on implementing POGIL covering the formation and effective management of student teams offering guidance on the selection and writing of POGIL activities as well as on facilitation teaching large classes and assessment The book concludes with examples of implementation in STEM and non STEM disciplines as well as guidance on how to get started Appendices provide additional resources and information about The POGIL Project

Argumentation in Chemistry Education Sibel Erduran, 2022-06-29 Scientists use arguments to relate the evidence that they select from their investigations and to justify the claims that they make about their observations This book brings together leading researchers to draw attention to research policy and practice around the inclusion of argumentation in chemistry education

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approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more *Organic Chemistry* Suzanne M. Ruder, The POGIL Project, 2015-12-29 ORGANIC CHEMISTRY **Nuts and Bolts of Chemical Education Research** Diane M. Bunce, Renée S. Cole, 2008 The purpose of this book is to address the key elements of planning chemical education research projects and educational outreach evaluation components of science grants from a pragmatic point of view *Advances in Teaching Physical Chemistry* Mark David Ellison, 2008 This book brings together the latest perspectives and ideas on teaching modern physical chemistry It includes perspectives from experienced and well known physical chemists a thorough review of the education literature pertaining to physical chemistry a thorough review of advances in undergraduate laboratory experiments from the past decade in depth descriptions of using computers to aid student learning and innovative ideas for teaching the fundamentals of physical chemistry This book will provide valuable insight and information to all teachers of physical chemistry **Chemists' Guide to Effective Teaching** Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 For courses in Methods of Teaching Chemistry Useful for new professors chemical educators or students learning to teach chemistry Intended for anyone who teaches chemistry or is learning to teach it this book examines applications of learning theories presenting actual techniques and practices that respected professors have used to implement and achieve their goals Each chapter is written by a chemist who has expertise in the area and who has experience in applying those ideas in their classrooms This book is a part of the Prentice Hall Series in Educational Innovation for Chemistry **Equilibrium** , 1991 **Equilibrium** Thomas R. Blackburn, 1969 **Chemical Equilibrium** William Guenther, 1975-12 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of

investigation *Equilibrium Calculations* John T. Donoghue,1971 **Equilibrium Constants of Liquid-liquid Distribution Reactions** International union of pure and applied chemistry. Commission on equilibrium data, **Equilibrium** Michell J. Sienko,1964 *The Nature of Equilibrium* Lifeliqe,2019 This lesson plan covers the nature of reversible reactions chemical equilibrium and how to calculate and use equilibrium constants **Chemical Equilibrium In a Nutshell** Brian Coleman,2018-06-25 Understanding the math and minutiae of chemical equilibrium can be a tall task for anyone so why not enlist the help of a scientific squirrel to guide you on your journey Join Dr Wash as we dabble in equilibrium constants and other tools needed to predict chemical processes This book focuses on introductory concepts at the high school and early university level focusing on identifying equilibrium calculating K and Q discussing Le Chatelier s principle and tying equilibrium with the field of Thermodynamics Full of step by step instructions and practice questions this book aims to simplify one of the more complex topics found within the field of chemistry **Chemical Equilibrium** William B. Guenther,1975 *Chemical equilibrium* , Chemical Equilibrium Robert Blanchard Fischer,Dennis G. Peters,1970 **Solution Equilibria** F. R. Hartley,Christopher Burgess,R. M. Alcock,1980 **Unified Equilibrium Calculations** William Benton Guenther,1991-08-19 A new approach to the academic treatment of solution equilibria is presented The author unifies homonuclear equilibrium calculations in one concept The alpha species fraction and bound proton and bound ligand ratio as a function of a single master variable the unbound H or L yield complete balances A single logic is maintained for all cases by equating the chemical binding expressed as an equilibrium condition and as a material balance condition

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