

Step I: Calculation of K_p for the reaction

Let the total mass of the gaseous mixture = 100 g

Mass of CO in the mixture = 90.55 g

Mass of CO₂ in the mixture = (100 - 90.55) = 9.45 g

$$\text{No. of moles of CO} = \frac{90.55 \text{ g}}{(28 \text{ g mol}^{-1})} = 3.234 \text{ mol}$$

$$\text{No. of moles of CO}_2 = \frac{9.45}{(44 \text{ g mol}^{-1})} = 0.215 \text{ mol}$$

$$p_{\text{CO}} \text{ in the mixture} = \frac{(3.234 \text{ mol})}{(3.234 + 0.215)} \times 1 \text{ atm} = \frac{(3.234 \text{ mol})}{(3.449 \text{ mol})} \times 1 \text{ atm} = 0.938 \text{ atm}$$

$$p_{\text{CO}_2} \text{ in the mixture} = \frac{(0.215 \text{ mol})}{(3.449 \text{ mol})} \times 1 \text{ atm} = 0.062 \text{ atm}$$



$$K_p = \frac{p^2_{\text{CO}}}{p_{\text{CO}_2}} = \frac{(0.938 \text{ atm})^2}{(0.062 \text{ atm})} = 14.19 \text{ atm}$$

Step II. Calculation of K_c for the reaction.

$$K_c = \frac{K_p}{(RT)^{\Delta n_g}}$$

$$K_p = 14.19 \text{ atm,}$$

$$R = 0.0821 \text{ L atm K}^{-1} \text{ mol}^{-1}, T = 1127 \text{ K}; \Delta n_g = 2 - 1 = 1$$

$$K_c = \frac{(14.19 \text{ atm})}{(0.0821 \text{ L atm K}^{-1} \text{ mol}^{-1}) \times (1127 \text{ K})^1} = 6.46$$

Equilibrium Chemistry Problems

Graeme E. Batley



Equilibrium Chemistry Problems:

Equilibrium Calculations John T. Donoghue, 1971 **Freshman Chemistry Problems and how to Solve Them** Michell J. Sienko, 1964 **A Problem-Solving Approach to Aquatic Chemistry** James N. Jensen, 2023-01-05

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F. Bleam, 2016-11-30 *Soil and Environmental Chemistry* Second Edition presents key aspects of soil chemistry in environmental science including dose responses risk characterization and practical applications of calculations using spreadsheets The book offers a holistic practical approach to the application of environmental chemistry to soil science and is designed to equip the reader with the chemistry knowledge and problem solving skills necessary to validate and interpret

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