

MCQ on Periodicity of Elements

1. The modern periodic table based on - **Atomic number**
2. On the basis of atomic radii which one is correct? - **$F < O < N$**
3. Group 18th elements is also known as - **Noble gases**
4. The most reactive element is - **Fluorine**
5. Transition metals belongs to - **Group number 3 to 12**
6. Which one has two completely filled shell? - **Neon**
7. Moving down the group of the periodic table, the factor that remains constant is - **Valence electrons**
8. Select the one that belongs to the same group as that of fluorine - **Astatine**
9. Which of the following is correct order of metallic character following elements is - **$K > Mg > Al > B$**
10. The element with atomic number 33 is from - **Group 15**
11. In the long form of the periodic table, the transition metals belong to - **d-block**
12. The correct order of their non-metallic character is - **$F > N > C > B > Si$**
13. The number of elements present in the 4th period - **18**
14. What would be the symbol for the element with atomic number 118? - **uuo**
15. Which of the following is not an actinoid? - **Terbium ($Z = 65$)**

Chemistry Objective Questions

Mcq For Inorganic Chemistry

Mr. Ajeet Singh



Mcq For Inorganic Chemistry:

MCQs Chemistry V. K. Ahluwalia, 2013-12-30 This book has been divided in 37 chapters under three parts Inorganic Chemistry Organic Chemistry and Physical Chemistry for convenient understanding It also includes solved model test papers of the previous three years of AIIMS CBSE PMT CPMT UP to enable students to develop the skills of problem solving and time management essential for any entrance examination In addition to providing answers to all the questions detailed explanatory notes to selected difficult questions have also been provided to justify the answer A separate section of Assertions and Reasons is also given at the end of each chapter Exhaustive Question Bank Explanatory Notes and Hints Assertions Reasons Includes Pre solved papers of five years Models Test Papers of AIIMS CBSE PMT CPMT [Class 11-12 Chemistry MCQ \(Multiple Choice Questions\)](#) Arshad Iqbal, 2019-05-17 The Class 11 12 Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF College Chemistry MCQ PDF Download Quiz Questions Chapter 1 6 Practice Tests with Answer Key 11th 12th Grade Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs Class 11 12 Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests Class 11 12 Chemistry MCQ PDF book helps to practice test questions from exam prep notes The Class 11 12 Chemistry MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs Class 11 12 Chemistry Multiple Choice Questions and Answers MCQs PDF Free download chapter 1 a book covers solved quiz questions and answers on chapters atomic structure basic chemistry chemical bonding chemistry experimental techniques gases liquids and solids tests for college and university revision guide Class 11 12 Chemistry Quiz Questions and Answers PDF free download eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The book Grade 11 12 Chemistry MCQs Chapter 1 6 PDF includes college question papers to review practice tests for exams Class 11 12 Chemistry Multiple Choice Questions MCQ with Answers PDF digital edition eBook a study guide with textbook chapters tests for NEET MCAT GRE GMAT SAT ACT competitive exam College Chemistry Mock Tests Chapter 1 6 eBook covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Atomic Structure MCQ Chapter 2 Basic Chemistry MCQ Chapter 3 Chemical Bonding MCQ Chapter 4 Experimental Techniques MCQ Chapter 5 Gases MCQ Chapter 6 Liquids and Solids MCQ The Atomic Structure MCQ PDF e Book Chapter 1 practice test to solve MCQ questions on Atoms atomic spectrum atomic absorption spectrum atomic emission spectrum molecules azimuthal quantum number Bohr s model Bohr s atomic model defects charge to mass ratio of electron discovery of electron discovery of neutron discovery of proton dual nature of matter electron charge electron distribution electron radius and energy derivation electron velocity electronic configuration of elements energy of revolving electron fundamental particles Heisenberg s uncertainty principle hydrogen spectrum magnetic quantum number mass of electron metallic crystals properties Moseley law neutron properties orbital concept photons wave number Planck s quantum theory properties of

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