

Quiz 132

Assessment

# Ions in Aqueous Solutions and Colligative Properties

## Section Quiz: Colligative Properties of Solutions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

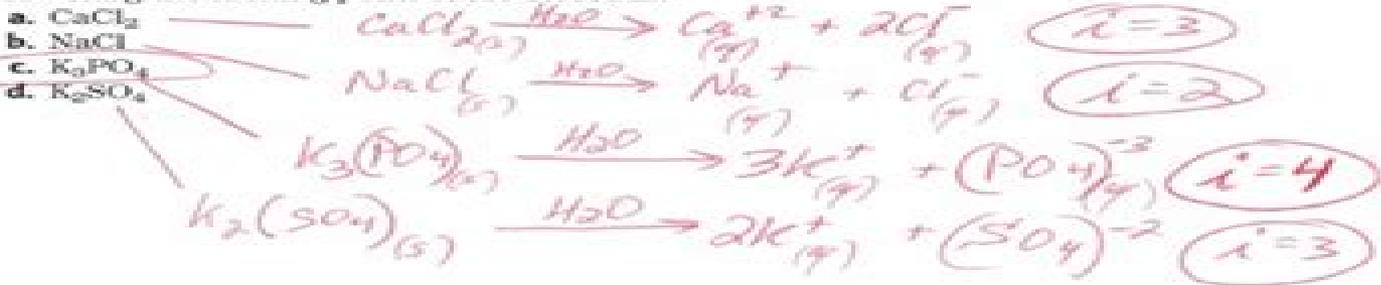
- B 1. Why is freezing-point depression a colligative property?
- a. It is inversely proportional to the number of particles in a solution.
  - b. It is directly proportional to the number of particles in a solution.
  - c. It depends on the properties of an electrolyte in a solvent.
  - d. None of the above

- B 2. Compared with a 0.01 M  $C_6H_{12}O_6$  solution, a 0.01 M KCl solution has
- a. the same freezing-point depression.
  - b. about twice the freezing-point depression.
  - c. the same freezing-point elevation.
  - d. about six times the freezing-point elevation.
- Handwritten notes:  $C_6H_{12}O_6$  (i=1), KCl (i=2)*

- D 3. Compared with a 1.00 M NaI solution, a 1.00 M  $Na_2SO_4$  solution has
- a. the same boiling-point elevation.
  - b. about twice the boiling-point elevation.
  - c. a boiling-point elevation about two-thirds as high.
  - d. a boiling-point elevation about 1.5 times as high.
- Handwritten notes: NaI i=2,  $Na_2(SO_4)$  i=3*

- B 4. When a nonvolatile solute dissolves in a solvent, the vapor pressure of the solvent
- a. increases.
  - b. decreases.
  - c. stays the same.
  - d. changes depending on the solvent used.
- Handwritten note: Vapor Pressure Lowering (Colligative Property)*

- C 5. Which of the following compounds would be most effective in lowering the melting point of ice on roads?



# Math Ch 13 Quiz 2 Answer Key

**Arshad Iqbal**



## **Math Ch 13 Quiz 2 Answer Key:**

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## **Table of Contents Math Ch 13 Quiz 2 Answer Key**

1. Understanding the eBook Math Ch 13 Quiz 2 Answer Key
  - The Rise of Digital Reading Math Ch 13 Quiz 2 Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Math Ch 13 Quiz 2 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Math Ch 13 Quiz 2 Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Math Ch 13 Quiz 2 Answer Key
  - Personalized Recommendations
  - Math Ch 13 Quiz 2 Answer Key User Reviews and Ratings
  - Math Ch 13 Quiz 2 Answer Key and Bestseller Lists
5. Accessing Math Ch 13 Quiz 2 Answer Key Free and Paid eBooks
  - Math Ch 13 Quiz 2 Answer Key Public Domain eBooks
  - Math Ch 13 Quiz 2 Answer Key eBook Subscription Services

- Math Ch 13 Quiz 2 Answer Key Budget-Friendly Options
- 6. Navigating Math Ch 13 Quiz 2 Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Math Ch 13 Quiz 2 Answer Key Compatibility with Devices
  - Math Ch 13 Quiz 2 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Math Ch 13 Quiz 2 Answer Key
  - Highlighting and Note-Taking Math Ch 13 Quiz 2 Answer Key
  - Interactive Elements Math Ch 13 Quiz 2 Answer Key
- 8. Staying Engaged with Math Ch 13 Quiz 2 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Math Ch 13 Quiz 2 Answer Key
- 9. Balancing eBooks and Physical Books Math Ch 13 Quiz 2 Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Math Ch 13 Quiz 2 Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Math Ch 13 Quiz 2 Answer Key
  - Setting Reading Goals Math Ch 13 Quiz 2 Answer Key
  - Carving Out Dedicated Reading Time
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