

Chapter 9 Review

- 1) What does stoichiometry mean? "measuring elements"
- 2) Before performing a stoichiometry problem, you must first do what with the chemical equation? **balance it**
- 3) The coefficients in a chemical equation represent what? **moles**
- 4) In a chemical equation, what relationships are shown? **molar ratios**
- 5) Solve the following mole-mole conversions using this equation:



- a) If 3.5 moles of sodium bicarbonate are decomposed, how many moles of carbon dioxide are produced?

$$3.5 \text{ mol NaHCO}_3 \times \frac{1 \text{ mol CO}_2}{2 \text{ mol NaHCO}_3} = 1.8 \text{ mol CO}_2$$

- b) If 1.29 moles of water are produced by the decomposition of sodium bicarbonate, how many moles of sodium carbonate are produced?

$$1.29 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{1 \text{ mol H}_2\text{O}} = 1.29 \text{ mol Na}_2\text{CO}_3$$

- c) How many moles of sodium bicarbonate are required to produce 19.5 moles of carbon dioxide?

$$19.5 \text{ mol CO}_2 \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol CO}_2} = 39.0 \text{ mol NaHCO}_3$$

- 6) Use the equation in # 5 to solve the following mass-mass conversions.

- a) If given 10.5 grams of sodium bicarbonate, how many grams of sodium carbonate can be produced?

$$10.5 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{2 \text{ mol NaHCO}_3} \times \frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} = 6.62 \text{ g Na}_2\text{CO}_3$$

- b) How many grams of sodium bicarbonate would be required to produce 9.0 grams of water?

$$9.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol H}_2\text{O}} \times \frac{84.01 \text{ g NaHCO}_3}{1 \text{ mol NaHCO}_3} = 84 \text{ g NaHCO}_3$$

- 7) What conversion factor (mole ratio) would you use to solve for moles of sulfur formed from sulfur dioxide in the reaction



Introduction To Moles Chemistry Packet Answers

Thomas Griffiths



Introduction To Moles Chemistry Packet Answers:

An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life. It provides a major overview of topics typically found in first year chemistry courses in the USA. The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box. Numerous examples are presented in every chapter to aid students and provide helpful self learning tools. The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry.

Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A.

Bartlett, Judith G. Koch, 1992 *Introduction to Chemistry, Study Guide* T. R. Dickson, 1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives. Three themes are stressed throughout the text: developing chemical thinking and chemical vision and refining problem solving skills. Many chapters in this edition have been rewritten and rearranged to vitalize the topics and to include interesting examples, analogies and images.

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