

CHAPTER 6 KINETICS

Page 162,163 Ex 6.1

1. B
2. C
3. A
4. a) $= 0.005/120 = 4.16 \times 10^{-5} \text{ mol dm}^{-3} \text{ s}^{-1}$
 b) i) H_2O_2 is consumed at a rate of $4.16 \times 10^{-5} \text{ mol dm}^{-3} \text{ s}^{-1}$
 ii) Γ^- ions are consumed at a rate of $8.32 \times 10^{-5} \text{ mol dm}^{-3} \text{ s}^{-1}$
 iii) is double i), because 2 moles of Γ^- are consumed for every mole of H_2O_2
5. (a)



- b) The reaction rate is greatest at the start of the reaction. This can be seen by the fact that the gradient of the curve is greatest at the start of the reaction.
- c) Possibilities include:
 - Collecting the gas and recording its volume against time.
 - Recording the change in total mass of the beaker and contents against time.
 - Titrating samples from the reaction mixture with an aqueous alkali of known concentration.

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1. C
2. D
3. A
4. The determination of the reaction rate would be best determined using a spectrophotometer. Probably setting the wavelength to red (complimentary colour of green) or blue (complimentary colour of orange would yield acceptable results. (The best way would be to scan the complete spectrum of both dichromate(VI) and chromium(III) and then select the wavelength where they have the greatest difference in absorption.). Mix together aqueous dichromate(VI) and aqueous sulfur dioxide (suitable concentrations to give a reasonable rate of reaction would have to be determined by trial and error), then transfer the mixture rapidly to a cuvette in the spectrophotometer. Record the absorption of the solution at different times, or preferably continually using a link to a computer. Draw a graph of absorption against time. The relative rate can be determined from the gradient of the graph. If a concentration-absorption calibration has been carried out at the wavelength used, this can be converted to a concentration-time graph and the absolute rate calculated. The major precaution would be to measure the temperature and try to keep it constant.

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