

## NOTES - Unit 4: Bonding

### PART 1: Ionic and Covalent Bonding

#### CHEMICAL BONDS

- A chemical bond is an attraction between 2 atoms or ions.
- Bonding occurs because it \_\_\_\_\_ of the system.
- Three broad classifications (general rule of thumb)
  - Ionic ( \_\_\_\_\_ + \_\_\_\_\_ )
  - Covalent ( \_\_\_\_\_ + \_\_\_\_\_ )
  - Metallic ( \_\_\_\_\_ + \_\_\_\_\_ )
- \_\_\_\_\_: the power of an atom in a molecule to attract electrons to itself.
  - The greater an atom's electronegativity, the greater its ability to attract \_\_\_\_\_.
  - In a compound, the element with the greater electronegativity will be the more \_\_\_\_\_ species.
  - The element with \_\_\_\_\_ electronegativity is sometimes referred to as the more " \_\_\_\_\_ " element. This will become the more positive species in the bond.
- \_\_\_\_\_ bonds exist between atoms with low electronegativities & those with high electronegativities. In general, between metals and nonmetals.
- \_\_\_\_\_ bonds exist between dissimilar atoms with different electronegativities.
- \_\_\_\_\_ bonds exist between identical nonmetallic atoms or nonmetallic atoms with similar electronegativities.
- Atoms react with each other in chemical reactions in a quest to have complete outer electron energy levels like the \_\_\_\_\_.
  - Octet Rule: Atoms tend to gain, lose, or share electrons until they are surrounded by \_\_\_\_\_ valence electrons.
  - Duet Rule: \_\_\_\_\_ and \_\_\_\_\_ want to have 2 electrons. Recall that the first energy level only holds two electrons.

#### IONIC BONDING

- Different texts report different "cutoff" values, but according to IB, \_\_\_\_\_ for ionic bonds.
- The greater the difference in electronegativity between the atoms, the more \_\_\_\_\_ the bond.
- Ionic bonds form between metals on the \_\_\_\_\_ of the periodic table and elements in groups \_\_\_\_\_ (or polyatomic ions attracted to other ions)
  - Note: group 4 elements (C and Si) do not form ionic bonds. They tend to form giant molecular structures (network solids) or simple molecules.
- Example: NaCl
  - Sodium (Na) has 11 electrons (but only one valence electron): \_\_\_\_\_
  - By losing this electron Na becomes  $\text{Na}^+$ : \_\_\_\_\_ or \_\_\_\_\_ which has a full outer shell.
  - Chlorine (Cl) has 17 electrons (seven valence electrons): \_\_\_\_\_

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