

# The Boltzmann Factor and Partition Functions

## PROBLEMS AND SOLUTIONS

**3-1.** How would you describe an ensemble whose systems are one-liter containers of water at 25°C?

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An unlimited number of one-liter containers of water in an essentially infinite heat bath at 298 K.

**3-2.** Show that Equation 3.8 is equivalent to  $f(x+y) = f(x)f(y)$ . In this problem, we will prove that  $f(x) \propto e^{\alpha x}$ . First, take the logarithm of the above equation to obtain

$$\ln f(x+y) = \ln f(x) + \ln f(y)$$

Differentiate both sides with respect to  $x$  (keeping  $y$  fixed) to get

$$\begin{aligned} \left[ \frac{\partial \ln f(x+y)}{\partial x} \right]_y &= \frac{d \ln f(x+y)}{d(x+y)} \left[ \frac{\partial(x+y)}{\partial x} \right]_y = \frac{d \ln f(x+y)}{d(x+y)} \\ &= \frac{d \ln f(x)}{dx} \end{aligned}$$

Now differentiate with respect to  $y$  (keeping  $x$  fixed) and show that

$$\frac{d \ln f(x)}{dx} = \frac{d \ln f(y)}{dy}$$

For this relation to be true for all  $x$  and  $y$ , each side must equal a constant, say  $\alpha$ . Show that

$$f(x) \propto e^{\alpha x} \quad \text{and} \quad f(y) \propto e^{\alpha y}$$

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Let  $x = E_1 - E_2$  and  $y = E_2 - E_3$ . Then  $x+y = E_1 - E_3$ , and we can write Equation 3.8 as  $f(x+y) = f(x)f(y)$ . Taking the logarithm of this equation gives

$$\begin{aligned} \ln f(x+y) &= \ln [f(x)f(y)] \\ &= \ln f(x) + \ln f(y) \end{aligned}$$

We then differentiate with respect to  $x$  and find

$$\begin{aligned} \left[ \frac{\partial \ln f(x+y)}{\partial x} \right]_y &= \left[ \frac{\partial \ln f(x)}{\partial x} \right]_y + 0 \\ \frac{d \ln f(x+y)}{d(x+y)} \left[ \frac{\partial(x+y)}{\partial x} \right]_y &= \frac{d \ln f(x)}{dx} \\ \frac{d \ln f(x+y)}{d(x+y)} &= \frac{d \ln f(x)}{dx} \end{aligned} \tag{1}$$

# Introduction To Molecular Thermodynamics Solution Manual

**Clemens Wendtner**



## **Introduction To Molecular Thermodynamics Solution Manual:**

The book delves into Introduction To Molecular Thermodynamics Solution Manual. Introduction To Molecular Thermodynamics Solution Manual is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Introduction To Molecular Thermodynamics Solution Manual, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
    - Chapter 1: Introduction to Introduction To Molecular Thermodynamics Solution Manual
    - Chapter 2: Essential Elements of Introduction To Molecular Thermodynamics Solution Manual
    - Chapter 3: Introduction To Molecular Thermodynamics Solution Manual in Everyday Life
    - Chapter 4: Introduction To Molecular Thermodynamics Solution Manual in Specific Contexts
    - Chapter 5: Conclusion
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  4. In chapter 3, the author will examine the practical applications of Introduction To Molecular Thermodynamics Solution Manual in daily life. The third chapter will showcase real-world examples of how Introduction To Molecular Thermodynamics Solution Manual can be effectively utilized in everyday scenarios.
  5. In chapter 4, this book will scrutinize the relevance of Introduction To Molecular Thermodynamics Solution Manual in specific contexts. The fourth chapter will explore how Introduction To Molecular Thermodynamics Solution Manual is applied in specialized fields, such as education, business, and technology.
  6. In chapter 5, the author will draw a conclusion about Introduction To Molecular Thermodynamics Solution Manual. This chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Introduction To Molecular Thermodynamics Solution Manual.

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