CHAPTER 9 REVIEW

Stoichiometry

SECTION 1

SHORT ANSWER Answer the following questions in the space provide	Get lev	CONTRACT	CONTRACT	400	ALBERT W.	24.00	MINNEY.	1000	and still departs of	SE 10	ModRowwitch	Till the second	Aller Streen and Co.	建设设置 医多种性 医	S B B COURT III
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- . The coefficients in a chemical equation represent the
 - (a) masses in grams of all reactants and products.
 - (b) relative number of moles of reactants and products.
 - (c) number of atoms of each element in each compound in a reaction.
 - 6d) number of valence electrons involved in a reaction.
- Which of the following would not be studied within the topic of stoichiometry?
 - (a) the mole ratio of Al to Cl in the compound aluminum chloride
 - (b) the mass of carbon produced when a known mass of sucrose decomposes
 - (c) the number of moles of hydrogen that will react with a known quantity of oxygen
 - (d) the amount of energy required to break the ionic bonds in CaF₂
- A balanced chemical equation allows you to determine the
 - (a) mole ratio of any two substances in the reaction.
 - (b) energy released in the reaction.
 - (c) electron configuration of all elements in the reaction.
 - (d) reaction mechanism involved in the reaction.
- The relative number of moles of hydrogen to moles of oxygen that react to form water represents admir
 - (a) reaction sequence.
 - (b) bond energy:
 - (c) mode ratio.
 - (d) element proportion.
- Given the reaction represented by the following unbalanced equation: N₂O(g) + O₂(g) → NO₂(g)
 - a. Balance the equation.
 - b. What is the mole ratio of NO₂ to O₂?
 - c. If 20.0 mol of NO₂ form, how many moles of O₂ must have been coom summed P.
 - d. Twice as many moles of NO₂ form as moles of N₂O are consumed. True or Folloc?
 - Twice as many grams of NOs form as grams of NsO are consumed. Trusc on Faller?

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Marcel A. Müller

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