

**ENGAGE**

Why does the atomic number but not the mass number change in positron emission?

**SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission****TRY IT FIRST**

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

**SOLUTION GUIDE**

	Given	Need	Connect
<b>ANALYZE THE PROBLEM</b>	Mn-49, positron emission	balanced nuclear equation	mass number, atomic number of new nucleus

**STEP 1** Write the incomplete nuclear equation.



**STEP 2** Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = ? + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49 \text{ (mass number of new nucleus)}$$

**STEP 3** Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = ? + 1$$

$$25 - 1 = ?$$

$$25 - 1 = 24 \text{ (atomic number of new nucleus)}$$

**STEP 4** Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written  ${}^{49}_{24}\text{Cr}$ .

**STEP 5** Complete the nuclear equation.

**STUDY CHECK 5.4**

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

**ANSWER****TEST**

Try Practice Problems 5.15 to 5.18

# Introduction To Chemical Answers Mastering Chemistry

**Phil Reedy, Donald J. Wink, Sharon  
Fetzer-Gislason**



## **Introduction To Chemical Answers Mastering Chemistry:**

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*CLASS 12 MASTERING ORGANIC REACTIONS COMPREHENSIVE GUIDE TO ORGANIC CHEMISTRY REACTIONS* Priti Singhal, 2024-11-11 The primary objective of this book is to serve as a comprehensive guide for students educators and researchers by focusing on reaction mechanisms practical applications and problem solving techniques Organic chemistry is not just about memorizing equations and formulas it is about understanding how molecules interact change and influence each other under different conditions With that in mind this book emphasizes the logic and patterns behind organic reactions making it easier for readers to apply concepts across a variety of scenarios Each chapter of this book builds upon foundational knowledge ensuring a progressive learning experience From nucleophilic substitutions to pericyclic reactions and from oxidation reduction mechanisms to named reactions we cover both fundamental and advanced topics to cater to students at all levels Real world examples have been integrated throughout the chapters to show how organic reactions play essential roles in pharmaceuticals biochemistry agriculture and environmental science This approach bridges the gap between theory and practical applications helping readers appreciate the relevance of organic chemistry in daily life

**Chemistry John**

Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science Chemical Principles Peter Atkins, Loretta Jones, 2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry's frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding

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invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings      **Lab Experiments in Introductory Chemistry** Phil Reedy, Donald J. Wink, Sharon Fetzner-Gislason, 2003-03-21 The manual contains laboratory experiments written specifically for the prep chem lab as well as for the general chemistry course Available as a complete manual or custom published at <http://custompub.whfreeman.com>      **The Publishers' Trade List Annual**, 1978      **Introduction to General, Organic, and Biochemistry** Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

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