

Honors Chemistry  
Equilibrium Worksheet

**Equilibrium: ICE box practice problems**

1. In a 10.0L vessel at 1000K, 0.250 mol  $\text{SO}_2(\text{g})$  and 0.200 mol  $\text{O}_2(\text{g})$  react to form 0.162 mol  $\text{SO}_3(\text{g})$  at equilibrium. What is the  $K_c$  at 1000K for this reaction?

Balanced equation:  $\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \leftrightarrow \text{SO}_3(\text{g})$

|   | $\text{SO}_2$ | + | $\text{O}_2$ | $\leftrightarrow$ | $\text{SO}_3$ |
|---|---------------|---|--------------|-------------------|---------------|
| I |               |   |              |                   |               |
| C |               |   |              |                   |               |
| E |               |   |              |                   |               |

2. An aqueous solution of 0.00325M  $\text{Fe}^{3+}$  was reacted with 0.0020M  $\text{SCN}^-$  at 25°C. Calculate the value of the equilibrium constant for this reaction when the equilibrium concentration of  $\text{FeSCN}^{2+}$  was determined to be 0.000667M

|   | $\text{Fe}^{3+}$ | + | $\text{SCN}^-$ | $\leftrightarrow$ | $\text{FeSCN}^{2+}$ |
|---|------------------|---|----------------|-------------------|---------------------|
| I |                  |   |                |                   |                     |
| C |                  |   |                |                   |                     |
| E |                  |   |                |                   |                     |

3. At a certain temperature, 2.0 mole  $\text{NH}_3$  is introduced into a 2.0 L container, and the  $\text{NH}_3$  partially dissociates into by the following reaction:



At equilibrium, 1.0 mol  $\text{NH}_3$  remains, what is the value of  $K$  for this reaction?

|   | $2\text{NH}_3(\text{g})$ | $\leftrightarrow$ | $\text{N}_2(\text{g})$ | + | $3\text{H}_2(\text{g})$ |
|---|--------------------------|-------------------|------------------------|---|-------------------------|
| I |                          |                   |                        |   |                         |
| C |                          |                   |                        |   |                         |
| E |                          |                   |                        |   |                         |

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1. Understanding the eBook Ice Box Practice Problems
  - The Rise of Digital Reading Ice Box Practice Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying Ice Box Practice Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ice Box Practice Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ice Box Practice Problems

- Personalized Recommendations
- Ice Box Practice Problems User Reviews and Ratings
- Ice Box Practice Problems and Bestseller Lists
- 5. Accessing Ice Box Practice Problems Free and Paid eBooks
  - Ice Box Practice Problems Public Domain eBooks
  - Ice Box Practice Problems eBook Subscription Services
  - Ice Box Practice Problems Budget-Friendly Options
- 6. Navigating Ice Box Practice Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - Ice Box Practice Problems Compatibility with Devices
  - Ice Box Practice Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ice Box Practice Problems
  - Highlighting and Note-Taking Ice Box Practice Problems
  - Interactive Elements Ice Box Practice Problems
- 8. Staying Engaged with Ice Box Practice Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ice Box Practice Problems
- 9. Balancing eBooks and Physical Books Ice Box Practice Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ice Box Practice Problems
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ice Box Practice Problems
  - Setting Reading Goals Ice Box Practice Problems
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ice Box Practice Problems

- Fact-Checking eBook Content of Ice Box Practice Problems
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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