

 **K_{sp} and ICE Problems Explained**

Key: K_{sp} = solubility product
 ICE = Initial, Change, Equilibrium

1. Write the K_{sp} expression for each of the following salts and then calculate the K_{sp} value.



Initial	0.00M	0.00M	0.00M
Change	+0.001M	+0.002M	+0.002M
Equilibrium	0.001M	0.002M	0.002M

2. Write the K_{sp} expression for each of the following salts and then calculate the K_{sp} value.



Initial	0.00M	0.00M
Change	+0.0000001M	+0.0000001M
Equilibrium	0.0000001M	0.0000001M

3. Use your K_{sp} values to calculate the molar solubility of each of the following salts and then compare your results with the K_{sp} values listed below for each of the compounds.



If the initial concentration of PbCl_2 is 0.10M, then the K_{sp} value is 1.7E-5. This is the K_{sp} value for PbCl_2 .

Initial	0.10M	0.00M	0.00M	0.00M	0.00M
Change	-x	+x	+2x	+2x	+2x
Equilibrium	0.10 - x	x	2x	2x	2x

Ice Practice Problems Answer Key

H Kauffman

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