

Q# 7/ IB Chem/2013/s/tz1/Paper 2 Section B/Standard Level/q6

- (d) (i) electron pair acceptor; [1]
- (ii) NH_3 : Lewis base;
 BF_3 : Lewis acid;
 NH_3 has non-bonding/lone pair of electrons and BF_3 has only 3 pairs of electrons around B/incomplete octet; [3]
Reference to outer electrons of N and B is needed for the mark.

Q# 8/ IB Chem/2013/s/tz1/Paper 2 Section B/Standard Level/q6

- (c) (i)
$$\begin{array}{c} NH_3 \\ | \\ H-N-H \\ | \\ H \end{array} ; \quad \begin{array}{c} BF_3 \\ | \\ F-B-F \\ | \\ F \end{array} ;$$
 [2]
Accept any combination of lines, dots or crosses to represent electron pairs.
- (ii) sharing of electrons between atoms; [1]
- (iii) NH_3 : (trigonal/triangular) pyramidal;
 BF_3 : trigonal/triangular planar;
 NH_3 has 4 negative centres of charge/three bonding pairs and one lone pair and BF_3 has 3 negative centres of charge/three bonding pairs / *OWTTE*;
 (bond angles) 107° in NH_3 and 120° in BF_3 ; [4]
Accept 107.5° for NH_3 .
- (iv) BF_3 not polar as no net dipole moment / BF bond polarities cancel each other out / symmetrical distribution of charge;
 NH_3 polar as net dipole moment present / NH bond polarities do not cancel each other out / unsymmetrical distribution of charge; [2]
Accept suitable diagram showing dipole moments.
Do not accept electronegativities cancel out.

Q# 9/ IB Chem/2012/w/tz0/Paper 2 Section B/Standard Level/q5

- (b) (i) *Bronsted Lowry theory*: proton/ H^+ donor;
Lewis theory: electron-pair acceptor; [2]
- (iv) *Strong acid*: acid/electrolyte completely/100% dissociated/ionized in solution/water / *OWTTE* and *Weak acid*: acid/electrolyte partially dissociated/ionized in solution/water / *OWTTE*; [1]
- (vi) changes by $10^2/100$; [1]
Allow changes from 10^{-3} to 10^{-7} .
- (vii) $(NH_4)_2SO_4$; [1]

Q# 10/ IB Chem/2012/s/tz1/Paper 2 Section B/Standard Level/q6

- (d) *Strong acid*: acid/electrolyte (assumed to be almost) 100%/completely dissociated/ionized (in solution/water) / *OWTTE* and *Weak acid*: acid/electrolyte only partially/slightly dissociated/ionized (in solution/water) / *OWTTE*;
 $HBr(aq) \rightarrow H^+(aq) + Br^-(aq)$;
 $HF(aq) \rightleftharpoons H^+(aq) + F^-(aq)$; [3]



Ib Sl Chemistry 2013 Paper

AN Whitehead



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1. Understanding the eBook Ib Sl Chemistry 2013 Paper
 - The Rise of Digital Reading Ib Sl Chemistry 2013 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ib Sl Chemistry 2013 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ib Sl Chemistry 2013 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ib Sl Chemistry 2013 Paper
 - Personalized Recommendations
 - Ib Sl Chemistry 2013 Paper User Reviews and Ratings
 - Ib Sl Chemistry 2013 Paper and Bestseller Lists
5. Accessing Ib Sl Chemistry 2013 Paper Free and Paid eBooks
 - Ib Sl Chemistry 2013 Paper Public Domain eBooks
 - Ib Sl Chemistry 2013 Paper eBook Subscription Services
 - Ib Sl Chemistry 2013 Paper Budget-Friendly Options
6. Navigating Ib Sl Chemistry 2013 Paper eBook Formats

- ePub, PDF, MOBI, and More
- Ib Sl Chemistry 2013 Paper Compatibility with Devices
- Ib Sl Chemistry 2013 Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ib Sl Chemistry 2013 Paper
 - Highlighting and Note-Taking Ib Sl Chemistry 2013 Paper
 - Interactive Elements Ib Sl Chemistry 2013 Paper
- 8. Staying Engaged with Ib Sl Chemistry 2013 Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ib Sl Chemistry 2013 Paper
- 9. Balancing eBooks and Physical Books Ib Sl Chemistry 2013 Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ib Sl Chemistry 2013 Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ib Sl Chemistry 2013 Paper
 - Setting Reading Goals Ib Sl Chemistry 2013 Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ib Sl Chemistry 2013 Paper
 - Fact-Checking eBook Content of Ib Sl Chemistry 2013 Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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