

Chapter 7 Answers

Practice 7-1

1. 8 2. 8 3. 60 4. 240 5. 2,635 6. 1.92
7. 0.8125 8. 9 9. 1500 10. 8.75 11. 3.44
12. 7 13. 110 14. 30 15. 16 16. 119
17. 12 18. 48 19. 40 20. 56

Practice 7-2

1. $\sqrt{51}$ 2. $4\sqrt{7}$ 3. $2\sqrt{65}$ 4. $\sqrt{7}$
5. $2\sqrt{21}$ 6. $18\sqrt{2}$ 7. 210 in.^2 8. $24\sqrt{10} \text{ cm}^2$
9. $35\sqrt{114} \text{ ft}^2$ 10. 46 in. 11. 78 ft 12. 279 cm
13. 19 m 14. acute 15. right 16. obtuse
17. right 18. obtuse 19. acute

Practice 7-3

1. $x = 2; y = \sqrt{3}$ 2. $a = 4.5; b = 4.5\sqrt{3}$
3. $c = \frac{80\sqrt{3}}{3}; d = \frac{20\sqrt{3}}{3}$ 4. $8\sqrt{2}$ 5. $14\sqrt{2}$
6. $\frac{28\sqrt{3}}{3}$ 7. 2 8. $x = 15; y = 15\sqrt{3}$ 9. $3\sqrt{2}$
10. 42 cm 11. 5.9 in. 12. 10.4 ft, 12 ft 13. 170 in.^2
14. $w = \frac{10\sqrt{3}}{3}; x = 5; y = 5\sqrt{2}; z = \frac{2\sqrt{2}}{3}$ 15. $a = 4;$
 $b = 3$ 16. $p = 4\sqrt{3}; q = 4\sqrt{3}; r = 8; s = 4\sqrt{6}$

Practice 7-4

1. 48 cm^2 2. 784 in.^2 3. 11.4 ft^2 4. 90 m^2
5. 2400 in.^2 6. 374 ft^2 7. 160 cm^2 8. 176.25 in.^2
9. 54 ft^2 10. 42.5 square units 11. $96\sqrt{3}$ square units
12. $36\sqrt{3}$ square units 13. 45 cm^2 14. 226.2 in.^2
15. $49,500 \text{ m}^2$

Practice 7-5

1. $x = 7; y = 3.5\sqrt{3}$ 2. $a = 60; c = \frac{8\sqrt{3}}{3}; d = \frac{8\sqrt{3}}{3}$
3. $p = 4\sqrt{3}; q = 8; x = 30$ 4. $m\angle 1 = 45; m\angle 2 = 45;$
 $m\angle 3 = 90; m\angle 4 = 90$ 5. $m\angle 5 = 60; m\angle 6 = 30;$
 $m\angle 7 = 120$ 6. $m\angle 8 = 60; m\angle 9 = 30; m\angle 10 = 60$
7. $12\sqrt{3}$ 8. $36.75\sqrt{3}$ 9. $75\sqrt{3}$ 10. 120 in.^2
11. 137 in.^2 12. 97 in.^2

Practice 7-6

1. 32π 2. 16π 3. 7.8π 4. Samples: $\overline{DFA}, \overline{ABF}$
5. Samples: $\overline{FE}, \overline{FD}$ 6. Samples: $\overline{FEB}, \overline{EDA}$
7. Samples: $\overline{FE}, \overline{ED}$ 8. $\angle ACB$ 9. $\angle FCD$
10. 32 11. 43 12. 54 13. 39 14. 71
15. 90 16. 50 17. 220 18. 140 19. 220
20. 130 21. $6\pi \text{ in.}$ 22. $16\pi \text{ cm}$ 23. $\frac{3}{4}\pi \text{ m}$

Practice 7-7

1. 49π 2. 21.2 3. $\frac{92}{5}\pi$ 4. $\frac{92}{5}\pi - 21.2$ 5. $\frac{1}{3}\pi$
or $\frac{\pi}{3}$ 6. $\frac{1}{3}$ 7. $\frac{1}{10}\pi$ or $\frac{\pi}{10}$ 8. $\frac{\pi}{10} - \frac{1}{3}$ 9. 4π
10. $\frac{12}{5}\pi$ 11. $\frac{2}{3}\pi$ 12. $\frac{3}{2}\pi$ 13. $\frac{12}{5}\pi$ 14. 6π
15. $\frac{92}{5}\pi$ 16. $\frac{92}{5}\pi$ 17. $\frac{\pi}{5}$ 18. 3 19. 2
20. 10

Practice 7-8

1. 8.7% 2. 10.9% 3. 15.3% 4. $\frac{3}{50}$ 5. $\frac{7}{10}$
6. $\frac{3}{5}$ 7. $\frac{1}{10}$ 8. 1 9. $\frac{1}{3}$ 10. 22.2% 11. 4.0%
12. 37.5% 13. 20% 14. 50% 15. 40%
16. $33\frac{1}{3}\%$

Reteaching 7-1

1. Samples:



2. Samples:



3. 20 cm^2 4. 55 in.^2 5. 6 m 6. 8 ft 7. 4 in.

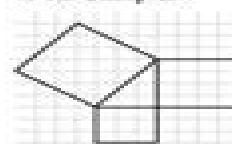
8. Samples:



9. 5 in. 10. 23 cm 11. 9 ft

Reteaching 7-2

1–2. Sample:



3. Sample: 9, 16, and 25 4. They are equal. 5. The sum of the squares of the two legs equals the square of the hypotenuse. 6. $4\sqrt{5}$ 7. $7\sqrt{2}$ 8. 8 9. 6

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Roman Wölfel



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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+ net ionic equation $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions

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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3 \text{Ba}^{2+} \text{aq} + 2 \text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web practice problems on net ionic equations page 2 of 3 answer key to practice problems on net ionic equations 1 molecular $\text{AgNO}_3 \text{aq} + \text{KCl} \text{aq} \rightarrow \text{AgCl} \text{s} + \text{KNO}_3 \text{aq}$ total ionic $\text{Ag}^+ \text{aq} + \text{NO}_3^- \text{aq} + \text{K}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s} + \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$ net ionic $\text{Ag}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s}$

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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web ionic equations and spectator ions in chemistry an ion is an atom or molecule that has a net electric charge due to the loss or gain of one or more electrons ions containing positive charge are cations and those containing negative charge are anions an ion with a net charge of zero is electrically neutral

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web apr 6 2022 spectator ion a spectator ion is an ion that exists as a reactant and a product in a chemical equation a spectator ion can therefore be observed in the reaction of aqueous solutions of sodium carbonate and copper ii sulfate but does not affect the equilibrium what are spectator ions

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web 18 terms lourdes l22 preview terms in this set 4 molecular equation showing the complete chemical formulas of reactions and products pb no3 2 aq 2kl aq pbi2 s 2kno3 aq complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the ag aq and cl aq ions become agcl s but the na aq ions and the no 3 aq ions stay as na aq ions and no 3 aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jul 26 2017 answer the spectator ions are explanation in the net ionic equations we are not include the spectator ions in the equations spectator ions the ions present on reactant and product side 16 used to grind chemicals to powder tmraor nda stlepe star 4 4 5 heart 181 verified verified answer food

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be 2cl aq cu 2 aq cucl 2 s 2cl aq cu 2 aq cucl 2 s these spectator ions are not actively involved in the reaction the existence of

spectator ions affects the debye length