Consider the isoelectronic species,  $Na^+$ ,  $Mg^{2+}$ ,  $F^-$ , and  $O^{2-}$ . The correct order of increasing length of their radii is  $1. F^- < O^{2-} < Mg^{2+} < Na^+ < 2. Mg^{2+} < Na^+ < F^- < O^{2-} < 3. O^{2-} < F^- < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < 4. O^{2-} < F^- < Mg^{2+} < Na^+ < Mg^{2+} < Mg^{2+$ 

Among the following elements, the one that is not an actinoid is: 1. Curium (Z=96) 2. Californium (Z=98)

The order of screening effect of electrons of s, p, d and f orbitals of a given shell of an atom on electrons in its outer shell is:

1. s > p > d > f

2. f > d > p > s 3. p < d < f < s 4. f > p > s > d

Uranium (Z=92).

Terbium (Z=65).

The first ionisation enthalpies of Na, Mg, Al, and Si are in the order of-1. Na < Al < Mg < Si 2. Na > Mg > Al > Si

2: Na > Mg > Al > Si 3. Na < Mg < Al < Si 4. Na > Mg > Al < Si

5 Among halogens, the correct order of amount of energy released in electron gain (electron gain enthalpy) is-

1, F > Cl > Br > I 2, F < Cl < Br < I 3, F < Cl > Br > I 4, F < Cl < Br < I

6 The period number in the long form of the periodic table is equal to:

 Magnetic quantum number of any element of the period.

Atomic number of any element of the period.

Maximum principal quantum number of any element of the period.

 Maximum azimuthal quantum number of any element of the period. The correct order of size of the given species is:

1. I > I' > I'

2.  $\Gamma^+ > \Gamma^- > \Gamma$ 

3.  $I > I^+ > I^-$ 

4.  $\Gamma > 1 > 1^+$ 

8 Electronic configuration of four elements A, B, C

and D are given below:

A.  $1s^2 2s^2 2p^6$ 

B.  $1s^2 2s^2 2p^4$ 

C.  $1s^2 2s^2 2p^6 3s^1$ 

D.  $1s^22s^22p^5$ 

The correct order of increasing tendency to gain electron is:

 $L, A \le C \le B \le D$ 

2. A < B < C < D

3.  $D \le B \le C \le A$ 

 $4. D \le A \le B \le C$ 

Match the correct ionization enthalpies and electron

gain enthalpies of the following elements.

	Elements		$\Delta H_1$	$\Delta H_2$	$\Delta_{\rm eg} H$
(i)	Most reactive non-metal	A.	419	3051	-48
(ii)	Most reactive metal	В.	1681	3374	-328
(iii)	Least reactive element	C.	738	1451	50
(iv)	Metal forming binary halide	D.	2372	5251	+48

#### Codes

	A	В	C	D
1.	iti	i.	in	iši
2.	i	111	101	iw
3.	1	iv	100	10
4.	ïv	i	iiii	111

# **Mcqs On Inorganic**

**Kumaresan Veliah** 



Whispering the Secrets of Language: An Emotional Quest through Mcqs On Inorganic

In a digitally-driven world where screens reign great and quick conversation drowns out the subtleties of language, the profound secrets and emotional nuances hidden within phrases usually go unheard. However, situated within the pages of **Mcqs On Inorganic** a charming literary prize pulsating with fresh emotions, lies an exceptional quest waiting to be undertaken. Penned by a talented wordsmith, that wonderful opus attracts viewers on an introspective trip, softly unraveling the veiled truths and profound impact resonating within the very fabric of every word. Within the psychological depths of this touching evaluation, we will embark upon a honest exploration of the book is core styles, dissect their captivating publishing fashion, and fail to the effective resonance it evokes strong within the recesses of readers hearts.

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