$$W + X \rightarrow Y + Z$$

26. Gases W and X react in a closed, rigid vessel to form gases Y and Z according to the equation above. The initial pressure of W is 1.20 atm and that of X is 1.60 atm. No Y(s) or Z(s) is initially present. The experiment is carried out at constant temperature. What is the partial pressure of $Z_{(g)}$ when the partial pressure of $W_{(g)}$ has decreased to 1.0 atm?

W+ x → y+ 2 30.20 atm (C)L0 atm (B)0.40 atm (D)1.2 atm (E)1.4 atm.

27 On a mountaintop, it is observed that water boils at 90 °C not at 100 °C as at sea level. This phenomenon occurs because on the mountaintop the

- (A) equilibrium water vapor pressure is higher due to the higher atmospheric pressure
- (B) equilibrium water vapor pressure is lower due to the higher atmospheric
- Cy equilibrium water vapor pressure equals the atmospheric pressure at a lower temperature
- (D) water molecules have a higher average kinetic energy due to the lower atmospheric pressure
- (E) water contains a greater concentration of dissolved gases

(C) 0.28 M

28 A 40.0 mL sample of 0.25 M KOH is added to 60.0 mL of 0.15 M Ba(OH). What is the molar concentration of OH (sag) in the resulting solution? (Assume that the volumes are additive.) .04 x . 25 = .01

(B) 0.19 M (A) 0.10 M(E) 0.55 M (D) 0.40 M

.06 x.15 = .004 .019/.1 = .19 NH₂NO₁₀₀ → N₂O₁₀₀ + 2 H₂O₁₀₀

29 A 0.03 mol sample of NH₄NO₂₀ is placed in a 1 L evacuated flask, which is then sealed and heated. The NH₄NO₂₀ decomposes completely according to the balanced equation above. The total pressure in the flask measured at 400 K is closest to which of the following? (The value of the gas constant, R, is 0.082 L atm mol 'K'.)

CA33 atm (B)1 atm (D)0.1 atm. (E)0.03 atm

(C)0.5 atm (1 XP)= .09(.012 Y400)

-1,323 kJ $C_2H_4(g) + 3 O_2(g) \rightarrow 2 CO_2(g) + 2 H_2O(g)$

 For the reaction of ethylene represented above, ΔH is -1,323 kJ. What is the value of ΔH if the combustion produced liquid water H₂O₍₀₎, rather than water vapor H₂O₍₀₎? $(\Delta H \text{ for the phase change } H_2O_{(2)} \rightarrow H_2O_{(2)} \text{ is } -44 \text{ kJ mol}^+.)$

(A) -1,235 kJ (C)-1,323 kJ (D)-1,367 kJ (E)-1,411 kJ

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 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - User-Friendly Interface
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
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 - Utilizing eBooks for Skill Development
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