

A2 Chemistry Topic 11 Rates

1. The kinetics of the hydrolysis of the halogenoalkane $R-CH_2Cl$ with aqueous sodium hydroxide (where R is an alkyl group) was studied at 50 °C. The following results were obtained:

Experiment	$[R-CH_2Cl]$	$[OH^-]$	Initial rate of $\text{dm}^{-3} \text{s}^{-1}$
1	0.050	0.10	4.0×10^{-3}
2	0.10	0.10	1.2×10^{-2}
3	0.10	0.20	1.6×10^{-2}

- (a) Deduce the order of reaction with respect to the halogenoalkane, $R-CH_2Cl$, and with respect to the hydroxide ion, OH^- , giving reasons for your answers.

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(3)

- (b) Hence write the rate equation for the reaction.

.....

(1)

- (c) Calculate the value of the rate constant with its units for this reaction at 50 °C.

(2)

2. 2-bromo-2-methylpropane reacts with aqueous sodium hydroxide as shown in the equation below.



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