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web the chemical equation used was $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product was determined to be ksp $\text{Ca}^{2+}\text{OH}^{-}$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error
accepted value for ksp for sodium hydroxide is 5×10^{-6}

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