

Activity A:

Subatomic particles

Get the Gizmo ready:

- Use the arrows to create an atom with two protons, two neutrons, and two electrons.
- Turn on **Show element name**.



Question: What are the properties of protons, neutrons, and electrons?

1. **Observe:** Turn on **Show element symbol** and **Element notation**. Three numbers surround the element symbol: the **mass number** (A), electrical charge (no number is displayed if the atom is neutral), and the **atomic number** (Z).
2. **Investigate:** Watch how the numbers change as you add or remove particles.



- A. Which number is equal to the number of protons in the atom?
- B. How can you calculate the number of neutrons (N) in an atom?
- C. Which particle (proton, neutron, or electron) has a
Positive charge?
Negative charge?

neutrons

by the number of protons

electron

proton

Electrical charge?

neutron

3. **Observe:** Turn on **Electron configuration**. An atom's electrons form a cloud around the nucleus. Electrons are arranged in shells. The number of shells is equal to the atomic number. The number of electrons in each shell is also equal to the atomic number. The number of electrons in each shell is also equal to the atomic number.

- A. What are the shells called?
- B. What are the shells called?
- C. What are the shells called?

Electron shells

Electron shells

Electron shells

4. **Observe:** Turn on **Electron configuration**. The following questions are asked:

- A. How many electrons are in a neutral atom of Helium?
- B. How many electrons are in a neutral atom of Hydrogen?
- C. What is the mass number of an atom with 2 protons and 2 neutrons?
- D. What is the mass number of an atom with 2 protons and 2 neutrons?
- E. How many electrons are in Helium?

2

1

4

4

2

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