Activity B:	Get the Gizmo ready:	0-000
Temperature and molecular motion	 Click Reset, and select the Micro view. Set Ice volume to 0 cm³. Set Add/remove heat energy to 0 J/s. 	6666

Question: Why do phase changes occur?

 Compare: Set the Water temperature to 0 °C and click Play. Observe the water molecules. Click Reset, set the Water temperature to 100 °C, and click Play again.

What do you notice?

Nothing differs, there is no difference.

- Observe: Click Reset. The mean molecular speed of the water molecules is displayed below the container. Set the Water temperature to 0 °C and Add/remove heat energy to 400 J/s. Click Play.
 - A. How does the mean speed of the water molecules change as they are heated?

Increases the hotter water gets

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, as the mean molecular speed varies as temperature rises.

3. Explain: How is temperature related to the motions of molecules?

The molecules move more quickly as the temperature rises.

4. Observe: Click Reset. Set the Water temperature to 20 °C and the Ice volume to 50 cm³. Set Add/remove heat energy to 0 J/s. Click Play. How do the molecules in the liquid interact with the molecules in the solid?

The bonds are broken when the molecules of the liquid collide with those of the solid.

 Observe: Click Reset. Set the Water temperature to 100 °C and the Ice volume to 50 cm³. Click Play. How does this situation compare to the previous one?

lce melts more quickly

Propose a theory: Based on what you have observed, explain why you think phase changes occur. If possible, discuss your theory with your classmates and teacher.

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Amaresh Chakrabarti, Udo Lindemann

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