

Total marks: 100

Total Time: 3 Hrs.

Attempt any five question. All questions carry equal marks.

- Q. No. 1 (a) What is zero-point energy? Write a detailed note on anharmonic oscillator and the selection rules governing the spectra. Explain the overtones from a diagram. (12)
(b) How many normal modes of vibration are possible for:
a. HBr b. SO₂ c. OCS d. C₆H₆ (4)
(c) What are aromatic compounds? (4)
- Q. No. 2 (a) Determine the formula of the following molecular/ions on the basis of MOT. Determine the bond order for each.
I. O₂⁻
II. N₂
III. Be₂²⁺
IV. Li₂ (6)
(b) Write a note on the relationship between occupancy, hybridization and geometry for the structures AB₃E, AB₂, AB₃, AB₄, AB₅, AB₆ type molecules. Give specific example for each. (12)
- Q. No. 3 (a) Give a general explanation for electrophilic substitution of benzene. Give chemical equation along with the mechanism showing all steps with reaction conditions for bromination, nitration, sulphonation and acylation. (10)
(b) How are human activities in urban areas responsible for air pollution? Explain with reference to smog and its chemistry. What measures should be taken to minimize air pollution? (10)
- Q. No. 4 (a) Debye-Hueckel Theory works under limiting conditions. Elaborate the statement with examples. Also give the significance of Debye Hueckel Theory. (14)
(2)
- Q. No. 5 (b) Calculate the ionic strength for a 0.1M solution of MgCl₂. (6)
(a) Write a note on standard electrode potential? How the direction of chemical reaction be established by a knowledge of it. Subsequently derive a relationship between ΔG° and K. (12)
(b) Use the following standard state cell potential to calculate the complex formation equilibrium constant K_f for the Zn(NH₃)₄²⁺ complex ion.
$$\text{Zn(NH}_3)_4^{2+} + 2e \leftrightarrow \text{Zn} + 4\text{NH}_3 \quad E_{\text{red}}^{\circ} = -1.04\text{V}$$
$$\text{Zn}^{2+} + 2e \leftrightarrow \text{Zn} \quad E_{\text{red}}^{\circ} = -0.762\text{V}$$
 (8)
- Q. No. 6 (a) Assign point groups to the following molecules. Write the symmetry for the point group with each molecule.
I. POCl₃
II. [PtCl₄]²⁻
III. H₂O
IV. SF₆ (16)
(b) Draw a multiplication table for the C_{3v} point group. (4)
- Q. No. 7 (a) Write a note on the effects of substitution groups on the rates of formation of carbocation in electrophilic aromatic substitution reactions when G releases electron and when it withdraws electron. Give specific examples of G in both cases. (8)
(b) Why a primary amine is more basic than ammonia which is more basic than a phenyl amine. Explain with respect to resonance structures and inductive effects. (6)
- Q. No. 8 (c) How is Global warming related to GHG? What is the basic condition for GHG to trap and re-emit radiation? (6)
(a) Synthesis of chiral compounds from achiral reactions always yield the racemic modification. Comment. Justify your answer with an example. (8)
(b) How amino acids behave as dipolar ions? Explain the Isoelectric point of amino acid. (8)
(c) What are zeolites and how can they be synthesized? (4)

F215 Past Papers 2014

Joacim Rocklöv



F215 Past Papers 2014:

F215 Past Papers 2014 Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**F215 Past Papers 2014**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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Table of Contents F215 Past Papers 2014

1. Understanding the eBook F215 Past Papers 2014
 - The Rise of Digital Reading F215 Past Papers 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying F215 Past Papers 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an F215 Past Papers 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from F215 Past Papers 2014
 - Personalized Recommendations
 - F215 Past Papers 2014 User Reviews and Ratings
 - F215 Past Papers 2014 and Bestseller Lists

5. Accessing F215 Past Papers 2014 Free and Paid eBooks
 - F215 Past Papers 2014 Public Domain eBooks
 - F215 Past Papers 2014 eBook Subscription Services
 - F215 Past Papers 2014 Budget-Friendly Options
6. Navigating F215 Past Papers 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - F215 Past Papers 2014 Compatibility with Devices
 - F215 Past Papers 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of F215 Past Papers 2014
 - Highlighting and Note-Taking F215 Past Papers 2014
 - Interactive Elements F215 Past Papers 2014
8. Staying Engaged with F215 Past Papers 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers F215 Past Papers 2014
9. Balancing eBooks and Physical Books F215 Past Papers 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection F215 Past Papers 2014
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine F215 Past Papers 2014
 - Setting Reading Goals F215 Past Papers 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of F215 Past Papers 2014
 - Fact-Checking eBook Content of F215 Past Papers 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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