



Exploring Equilibrium

It Works Both Ways

Introduction

The sound exact librium has two roots segret, meaning recast, and Strat, meaning weight or bulance. Our physical sense of equalitation—in the nestion of a service or the sering of a perchalance—suggests an equal bulance of appointing forces. How does this physical sense of equilibrium translate to chemical equilibrium? Explore the nature and consequences of equilibrium to the other order.

Concepts

Reversible machines.

Chemical equilibrium

· Commercion reaction

Activitude incitations

Background

Physical changes, such as nielting for or classifying suger, are often increcated by noting that these processes can be easily reversed. Some connon examples of chemical change, such as burning wood or spoiling, fixed, generally cannot be reversed. A closer look at chemical change, however, reveals that, many chemical reactions are also reversible.

Consider the following example of a reversible chemical reaction (Ecuation 1). At high pressures and in the presence of a catalyst, nitrogen and hydrogen react to form ammonia. If the temperature is high enough, however, ammonia decomposes to reform its constituent elements. The traction can go both way? This reversible reaction is represented symbolically using double amount.

Physical Science 2:

What happens when nitrogen and hydrogen are allowed to react? In a closed system, the concentrations of nitrogen and hydrogen will decrease, and the concentrations of anaroma will steadily increase as the maction process in the forward direction. Soon, however, the concentration of anaromic will be large enough that the reverse roution will begin to take place at a significant rate as well. Esembally, as the reaction occurs in both the forward and the research direction, the number of anaromic molecules being formed will become estad to the number of anaromic numbersheets being consumed. At this point, no further changes a will be observed in the overall concentrations of nitrogen, hydrogen and a nonconta. This is the point of chemical equilibrium.

Observined equilibrium is defined, as the state in which the rate of the forward reaction equals the rate of the reverse reaction and the concentrations of reactants and products remain constant, with time. Note that, this definition describes a dynamic picture of equilibrium. The reactions continue, but there is an equal balance of opposing reaction rates.

What happens when the southinions is disperbed? Any factor that changes the rate of the forward or the reverse reaction will change the associate of reactions and products that are present at equilibrium. Reaction conditions that are known to allow the rates of chemical reactions include the concentrations of reactions and the temperature. In this experiment, we will investigate how changes in traction conditions affect the associates of reactions are therefore a conditions of the associates of reactions are products present at equilibrium.

Experiment Overview

The purpose of this experiment is to explore the nature of chemical equilibrium and unitiaritify conditions that affect the position of equilibrium. Two different reversible reactions will be studied.

Reaction of iron(III) nitrate with potassium thiocyanate will be used to study complies interpretibrium. Iron(III) sons reset with thiocyanate torse to form PeSCN¹ complex ions (Equation 2). The offects of changing the concentrations of reactions and of changing the reaction temperature will be investigated.

Planation 2

The properties of an inclinator will be used to study and Autoropoliteiston. An inclinator is a dynthot one gain or lose hydrogen ions to force substances that have different orders. Equation 3 summarises the reservible reaction of the inclinator branches of green (HIrd). His represents an uncharged inclinator molecule and. In an indicator arisen formed after the molecule has lost a hydrogen ion. The color of the indicator in the presence of either reacts H. or CHI. ions (see Equation 4) will show have changing the concentration of a problem. Affects the equilibrium shows in Equation 3.

Cartine To.

Physics Phone 3

 $H(ac) + OH(ac) \rightarrow H_1O0$

Elpaintion 4

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Huangqi Zhang

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