



QUESTION: 10/10 (100% correct)

Time taken:

Score: 10/10 (100%)

10/10

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Your Result: correct for

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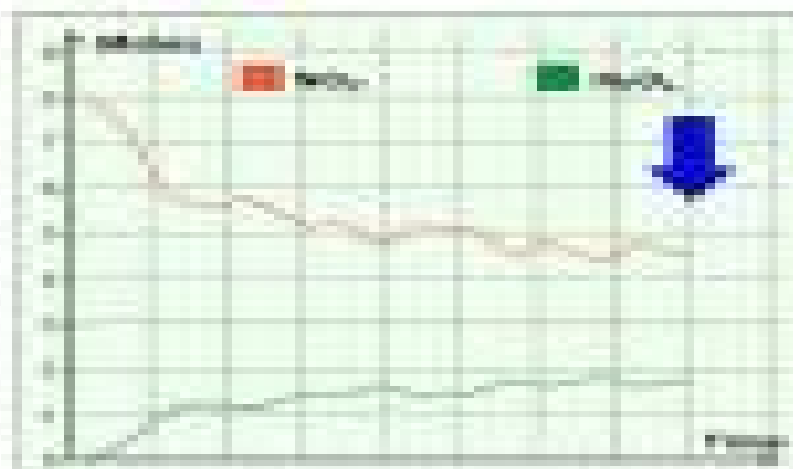
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Question 10 (100%)

A closed flask containing hydrogen gas at high pressure sits on a hot stove burner. Eventually, the hydrogen gas (H_2) escapes the flask. What is the thermodynamic quantity that is most directly responsible for this event?



- A. Entropy, because as the temperature rises, H_2O_2 decomposes.
- B. The reaction between oxygen and the flask causes the flask to break.
- C. The rate of hydrogen gas production is constant, so the rate of hydrogen gas production is constant.
- D. The activation energy of the reaction is low, so the reaction is fast.

Correct Answer: A. The graph shows the concentration of H_2O_2 decreasing over time, which is the most direct evidence for the reaction occurring.

Explanation: The graph shows the concentration of H_2O_2 (red line) decreasing over time, which is the most direct evidence for the reaction occurring. The green line shows the concentration of H_2O increasing over time, which is also evidence for the reaction occurring. The blue arrow points to the red line, indicating that the concentration of H_2O_2 is the most directly responsible for the event.

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Raffaela Di Napoli

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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