Name: __Chelsey Williams_____ Date: 3/31/22_____

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, formula mass, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

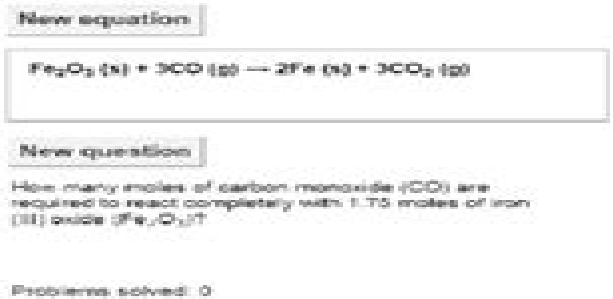
1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? 176 grams

   _____

2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? __15 sticks of butter

   _____

Gizmo Warm-up
Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The *Stoichiometry* Gizmo allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.



To begin, check that this equation is shown:

$$Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$$

1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

   Fe_2O_3 __1___ CO __3___ Fe __2___ CO_2 _3____

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

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RJ Alexander

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Table of Contents Explore Learning Gizmo Answer Key Stoichiometry

1. Understanding the eBook Explore Learning Gizmo Answer Key Stoichiometry
    - The Rise of Digital Reading Explore Learning Gizmo Answer Key Stoichiometry
    - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Gizmo Answer Key Stoichiometry
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Explore Learning Gizmo Answer Key Stoichiometry
    - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Gizmo Answer Key Stoichiometry
    - Personalized Recommendations
    - Explore Learning Gizmo Answer Key Stoichiometry User Reviews and Ratings
    - Explore Learning Gizmo Answer Key Stoichiometry and Bestseller Lists
5. Accessing Explore Learning Gizmo Answer Key Stoichiometry Free and Paid eBooks
    - Explore Learning Gizmo Answer Key Stoichiometry Public Domain eBooks
    - Explore Learning Gizmo Answer Key Stoichiometry eBook Subscription Services
    - Explore Learning Gizmo Answer Key Stoichiometry Budget-Friendly Options
6. Navigating Explore Learning Gizmo Answer Key Stoichiometry eBook Formats
    - ePub, PDF, MOBI, and More
    - Explore Learning Gizmo Answer Key Stoichiometry Compatibility with Devices
    - Explore Learning Gizmo Answer Key Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Explore Learning Gizmo Answer Key Stoichiometry
    - Highlighting and Note-Taking Explore Learning Gizmo Answer Key Stoichiometry
    - Interactive Elements Explore Learning Gizmo Answer Key Stoichiometry
8. Staying Engaged with Explore Learning Gizmo Answer Key Stoichiometry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Explore Learning Gizmo Answer Key Stoichiometry

9. Balancing eBooks and Physical Books Explore Learning Gizmo Answer Key Stoichiometry
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Explore Learning Gizmo Answer Key Stoichiometry
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine Explore Learning Gizmo Answer Key Stoichiometry
    - Setting Reading Goals Explore Learning Gizmo Answer Key Stoichiometry
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Explore Learning Gizmo Answer Key Stoichiometry
    - Fact-Checking eBook Content of Explore Learning Gizmo Answer Key Stoichiometry
    - Distinguishing Credible Sources
13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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