

82. Additional Questions

1. The percent by mass composition of a salt was found to be 76.58% potassium, 6.48% sulfur, and 16.94% oxygen. What is the empirical formula of this salt?

$$56.58g \text{ of K} \left(\frac{\text{molar K}}{39.098 \text{ g K}} \right) = 1.447 \text{ molar K} \quad \frac{1.447 \text{ molar K}}{0.723 \text{ molar S}} = 2$$

$$6.48g \text{ of S} \left(\frac{\text{molar S}}{32.065 \text{ g S}} \right) = 0.202 \text{ molar S} \quad \frac{0.202 \text{ molar S}}{0.202 \text{ molar S}} = 1$$

$$16.94g \text{ of O} \left(\frac{\text{molar O}}{16.000 \text{ g O}} \right) = 1.059 \text{ molar O} \quad \frac{1.059 \text{ molar O}}{0.202 \text{ molar S}} = 5$$

empirical formula = $K_2S_1O_5$

2. A compound consisting iron and sulfur was formed by combining 3.233 g of iron with 1.426 g of sulfur. What is the empirical formula of the compound?

$$3.233g \text{ of Fe} \left(\frac{\text{molar Fe}}{55.845 \text{ g Fe}} \right) = 0.05790 \text{ molar Fe} \quad \frac{0.05790 \text{ molar Fe}}{0.05790 \text{ molar Fe}} = 1 + 2 = 2$$

$$1.426g \text{ of S} \left(\frac{\text{molar S}}{32.065 \text{ g S}} \right) = 0.04447 \text{ molar S} \quad \frac{0.04447 \text{ molar S}}{0.05790 \text{ molar Fe}} = 0.768 + 2 = 3$$

empirical formula = Fe_2S_3

3. Propylene has a molar mass of 42.08 g/mole and is composed of 84.3% hydrogen and 15.7% carbon. What is the molecular formula of propylene?

$$14.3g \text{ of H} \left(\frac{\text{molar H}}{1.008 \text{ g H}} \right) = 14.15 \text{ molar H} \quad \frac{14.15 \text{ molar H}}{7.15 \text{ molar C}} = 2$$

$$15.7g \text{ of C} \left(\frac{\text{molar C}}{12.011 \text{ g C}} \right) = 1.307 \text{ molar C} \quad \frac{1.307 \text{ molar C}}{1.307 \text{ molar C}} = 1$$

empirical formula = CH_2

$$CH_2 = 12.011 + 2(1.008) = 14.027 \quad \frac{42.089}{14.027} = 3$$

molecular formula = C_3H_6

Magnesium Oxide Lab Answers

RJ Shavelson



Magnesium Oxide Lab Answers:

The book delves into Magnesium Oxide Lab Answers. Magnesium Oxide Lab Answers is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Magnesium Oxide Lab Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Magnesium Oxide Lab Answers
 - Chapter 2: Essential Elements of Magnesium Oxide Lab Answers
 - Chapter 3: Magnesium Oxide Lab Answers in Everyday Life
 - Chapter 4: Magnesium Oxide Lab Answers in Specific Contexts
 - Chapter 5: Conclusion
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 4. In chapter 3, this book will examine the practical applications of Magnesium Oxide Lab Answers in daily life. The third chapter will showcase real-world examples of how Magnesium Oxide Lab Answers can be effectively utilized in everyday scenarios.
 5. In chapter 4, the author will scrutinize the relevance of Magnesium Oxide Lab Answers in specific contexts. This chapter will explore how Magnesium Oxide Lab Answers is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Magnesium Oxide Lab Answers. This chapter will summarize the key points that have been discussed throughout the book.
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