Test A

Thermodynamics

MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question.

- If an ideal gas does positive work on its surroundings.
 - a. the temperature of the gas increases.
 - b. the volume of the gas increases.
 - e. the pressure of the gas increases.
 - d. the internal energy of the gas decreases.
- Sunlight strikes an ice cube at its melting point and causes it to melt. What is this process?
 - adiabatic process

isobaric process

b. isothermal process

- d. isovolumetric process
- 3. ______ Which of the following processes for an ideal gas system has an unchanging internal energy and a heat intake that corresponds to the value of the work done by the system?
 - isothermal process

e, adiabatic process

b. isobaric process

- d. isovolumetric process
- How is conservation of internal energy expressed for an isovolumetric system?
 - $\mathbf{a}_{-} Q = W = \Delta U = 0$; $U_i = U_f$
 - **b.** Q = 0; so $\Delta U = -W$
 - c. $\Delta T = 0$, so $\Delta U = Q W = 0$, so Q = W
 - **d.** $\Delta V = 0$, so $P\Delta V = 0$ and W = 0, so $\Delta U = Q$
- 5. _____ A 4 mol ideal gas system undergoes an adiabatic process in which it expands and does 20 J of work on its environment. What is its change in internal energy?
 - a. -20 J

e. 0 J

b. -5 J

- d. 20 J
- According to the second law of thermodynamics, the heat received by a heat engine operating in a complete cycle from a high-temperature reservoir
 - a. must be completely converted to work.
 - b. equals the entropy increase.
 - e. can be completely converted to internal energy.
 - d. cannot be completely converted to work.

Holt Physics Chapter 11 Test

Yemima Ben-Menahem, Meir Hemmo

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