

## EXPERIMENT 7 CHEMICAL KINETICS **SAMPLE REPORT**

### Special Notes

1. The report must be written in ink (blue pen), no pencil allowed. However, the graph can be drawn in pencil.
2. The layout of the report must be tidy (no white out) and legible in writing.  
(Work out all the solutions on some scrap papers first, then copy to the report sheets.)
3. Page 7 - 1
  - ① Record exact temperatures in °C to 1 decimal.
  - ② Both the "Time" in seconds and "Relative Rate" must be integers.
  - ③ Report the "Initial  $[I^-]$ ", "Initial  $[S_2O_8^{2-}]$ " and "Rate" in *3 s.f.*
4. Page 7 - 2
  - ① Show **all** calculations for:
    - (A) Concentration of initial  $[I^-]$ ,
    - (B) Concentration of initial  $[S_2O_8^{2-}]$ ,
    - (C) Rate of Reaction of  $S_2O_8^{2-}$ , and
    - (D) Relative Rates
  - ② Keep all calculations in *3 s.f.*. However, Relative Rate, Order of Reaction, and "m" and "n" must be reported as integers.
5. Page 7 - 3
  - ① For Question 1, the conclusion must include: **rate**  $\propto$   $[I^-]$ .
  - ② For Question 2, the conclusion must include: **rate**  $\propto$   $[S_2O_8^{2-}]$ .
  - ③ For Question 3, the rate law is: **rate**  $= k [I^-] [S_2O_8^{2-}]$
6. Page 7 - 4
  - ① Show calculations of rate constants, **k**, for all five reactions.
  - ② The **units** of rate constant, **k**, is  $M^{-1}s^{-1}$ .
  - ③ The **accuracy** of **k** is between  $3.30 \times 10^{-3}$  to  $3.90 \times 10^{-3} M^{-1}s^{-1}$  at 22-23°C.
  - ④ The **precision** of **k**  $\pm 0.30 \times 10^{-3} M^{-1}s^{-1}$  between highest and lowest **k**.

# Kinetics Of A Chemical Reaction Lab Report

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## **Table of Contents Kinetics Of A Chemical Reaction Lab Report**

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  - The Rise of Digital Reading Kinetics Of A Chemical Reaction Lab Report
  - Advantages of eBooks Over Traditional Books
2. Identifying Kinetics Of A Chemical Reaction Lab Report
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Kinetics Of A Chemical Reaction Lab Report
  - User-Friendly Interface
4. Exploring eBook Recommendations from Kinetics Of A Chemical Reaction Lab Report
  - Personalized Recommendations
  - Kinetics Of A Chemical Reaction Lab Report User Reviews and Ratings
  - Kinetics Of A Chemical Reaction Lab Report and Bestseller Lists
5. Accessing Kinetics Of A Chemical Reaction Lab Report Free and Paid eBooks

- Kinetics Of A Chemical Reaction Lab Report Public Domain eBooks
- Kinetics Of A Chemical Reaction Lab Report eBook Subscription Services
- Kinetics Of A Chemical Reaction Lab Report Budget-Friendly Options
- 6. Navigating Kinetics Of A Chemical Reaction Lab Report eBook Formats
  - ePub, PDF, MOBI, and More
  - Kinetics Of A Chemical Reaction Lab Report Compatibility with Devices
  - Kinetics Of A Chemical Reaction Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Kinetics Of A Chemical Reaction Lab Report
  - Highlighting and Note-Taking Kinetics Of A Chemical Reaction Lab Report
  - Interactive Elements Kinetics Of A Chemical Reaction Lab Report
- 8. Staying Engaged with Kinetics Of A Chemical Reaction Lab Report
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Kinetics Of A Chemical Reaction Lab Report
- 9. Balancing eBooks and Physical Books Kinetics Of A Chemical Reaction Lab Report
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Kinetics Of A Chemical Reaction Lab Report
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Kinetics Of A Chemical Reaction Lab Report
  - Setting Reading Goals Kinetics Of A Chemical Reaction Lab Report
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Kinetics Of A Chemical Reaction Lab Report
  - Fact-Checking eBook Content of Kinetics Of A Chemical Reaction Lab Report
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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