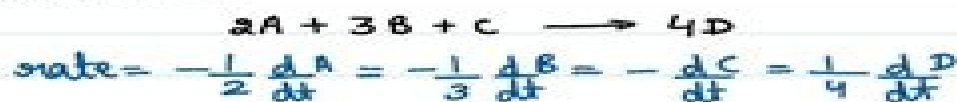


Chemical Kinetics

* rate of reaction



* $\text{rate} = k_n [A]^n$; n = order of reaction

$$\frac{x_1}{x_2} = \left(\frac{a_1}{a_2}\right)^n \Rightarrow \log \frac{x_1}{x_2} = n \log \frac{a_1}{a_2} \Rightarrow n = \frac{\log \frac{x_1}{x_2}}{\log \frac{a_1}{a_2}}$$

* zero order reaction:

$$a_x = a_0 - k_0 t$$

$$t_{1/2} = \frac{a_0}{2k_0}$$



* First order reaction

$$\ln \frac{a}{a-x} = k_1 t \quad t_{1/2} = \frac{0.693}{k_1}$$

$$k_1 t = \ln \frac{V_{\text{max}}}{V_{\text{max}} - V_t} \quad \text{For gas phase reactions}$$

* Higher order reactions

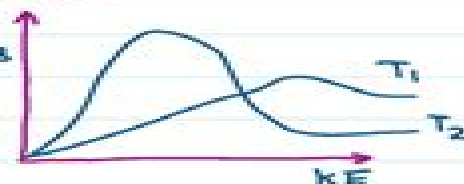
$$k_n t = \frac{1}{n-1} \left[\frac{1}{[A_t]^{n-1}} - \frac{1}{[A_0]^{n-1}} \right]$$

* Temperature coefficient

$$\frac{k_{T_1}}{k_{T_2}} = 2-3 = \phi \quad \text{when } T_1 = T_2 + 10^\circ\text{C}$$

maxwell distribution curve

no of molecules



* Arrhenius Equation

$$k = A e^{-E_a/RT} \Rightarrow \log k = \log A - \frac{E_a}{2.303RT}$$

$$\log \frac{k_1}{k_2} = \frac{E_a}{2.303R} \left[\frac{1}{T_2} - \frac{1}{T_1} \right]$$

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