

Three acid-base equilibria are shown in the following diagram:

acetic acid, a 100 mole % solution of acetic acid is placed in a 100 L container. The following reaction occurs:  $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$  in the container. The container is sealed and the concentration of the products is measured. The graph below shows:



1. Determine, based on the graph above, the **concentration** of the acetic acid at equilibrium and the products (acetic acid + water) at equilibrium.
2. Determine the change in the concentration of  $\text{CH}_3\text{COOH}$  from 0 to 10 seconds.
3. Determine, based on the following graph, the concentration of the products at equilibrium.

Use the following equation:

$$\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$$

Explain, in terms of energy, why the concentration of the products is higher than the concentration of the reactants.

4. Determine, based on the following graph, the concentration of the products at equilibrium.

A 100 mole % solution of acetic acid is placed in a 100 L container. The following reaction occurs:  $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$  in the container. The container is sealed and the concentration of the products is measured. The graph below shows:

Use the equation:

$$\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$$

Explain, in terms of energy, why the concentration of the products is higher than the concentration of the reactants.

Three different acid-base equilibria are shown in the following diagram. The concentration of the products is higher than the concentration of the reactants. The concentration of the products is higher than the concentration of the reactants.

# Kinetics And Equilibrium Test Answers

**JR Anderson**



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