

Reaction Coordinate

Which statement correctly describes the energy changes that occur in the forward reaction? The activation energy is 10, kJ and the

The activation energy is 10. kJ and the reaction is exothermic.

reaction is endothermic.

- (1) The activation energy is 50, kJ and the reaction is endothermic. (D) The activation energy is 50, kJ and the
- reaction is exothernole. 9. Given the reaction at equilibrium:

$$X_2(g) + 2 Y_2(g) \leftrightarrow 2 XY_2(g) + 80 keal$$

The equilibrium point will shift to the right if the pressure is

A) increased and the temperature is increased. increased and the temperature in decreased.



- -285.8 kg
- B) +285.8 kJ Do +571.6 kJ
- C) -571.6 kJ
- 11. As the number of effective collisions between reacting particles increases, the rate of reaction
 - A) decreases **ECCENSES** C) renains the same
 - 12. Which statement most be true when
 - solution equilibrium occum? A) The solution is at STP.
- The solution is supersuturated. The concentration of the solution remains constant.
- D) The masses of the dissolved solute and the undissolved solute are equal.
- 13. Given the reaction at equilibrium:

$$A(y) + B(y) \leftrightarrow C(y) + D(y)$$

The addition of a catalyst will A) shift the equilibrium to the right

- B) shift the equilibrium to the left. increase the rate of forward and reverse reactions equally
- have no effect on the forward or ritreme mactions
- 14. Which equation shows an increase in entropy?
 - A) CO₂(g) → CO₂(s) $CO_2(E) \rightarrow CO_2(e)$

Kinetics And Equilibrium Practice Test

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