

Completion

Complete each sentence or statement.

12. An electron that occupies the outermost energy level of an atom is known as a(n) Valence electron.
13. An atom or group of atoms that has an electrical charge because it has either lost or gained one or more electrons is a(n) ion.
14. An ion that has a negative charge is a(n) anion.
15. The ion formed by an atom of a metal is a(n) cation.
16. The ion formed by an atom of a nonmetal element is a(n) anion.
17. A compound resulting from the formation of an ionic bond between a cation and an anion is a(n) salt.
18. A repetitive geometric arrangement of ions, atoms, or molecules that forms a crystal structure is called the crystal lattice.
19. The name of the ion O^{2-} is the oxide ion.
20. The name of the ion Cu^+ is the copper(I) ion.
21. The chemical formula for the compound strontium sulfide, which contains Sr^{2+} and S^{2-} ions, is SrS .

Word Bank:

Valence	Crystal lattice	Cation	Anion	Copper(I)
ion	Salt	Oxide	Anion	SrS

Write the formula for the following ionic compounds:

1. Sodium chloride: $NaCl$
2. Calcium oxide: CaO
3. Iron (II) fluoride: FeF_2
4. Copper (I) oxide: Cu_2O
5. Aluminum sulfate: $Al_2(SO_4)_3$
6. Silver chloride: $AgCl$
7. Potassium carbonate: K_2CO_3
8. Tin (IV) fluoride: SnF_4
9. Chromium (III) oxide: Cr_2O_3
10. Sodium sulfide: Na_2S

Write the names for the following ionic compounds:

1. $NaNO_3$ Sodium Nitrate
2. Li_2O lithium oxide
3. $CuCO_3$ copper(II) carbonate
4. Fe_2O_3 Iron(III) oxide
5. $Ca_3(PO_4)_2$ calcium phosphate
6. SnO tin(II) oxide
7. $MgCO_3$ magnesium carbonate
8. $MnCO_3$ manganese(II) carbonate
9. Ag_2O silver oxide
10. $TiCl_3$ titanium(III) chloride

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we proceed as follows log x 2 0258 3 3 2 0258 3 9742 3 9742 therefore x antilog 3 9742 now from antilog table we get the

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698970004336019 5 now let's try it with a larger number

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consider the mantissa 5463

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