

# PRE-ALGEBRA FUN

## ANSWER KEY

**Direction:** Solve for the missing variable for each of the following equations.

$1.1 + y = 10y = 9$

$2.y + 9 = 14y = 5$

$3.9 + y = 15y = 6$

$4.y + 6 = 8y = 2$

$5.4 + y = 8y = 4$

$6.9 + y = 11y = 2$

$7.7 + y = 11y = 4$

$8.3 + y = 5y = 2$

$9.y + 2 = 3y = 1$

$10.4 + y = 13y = 9$

$11.4 + y = 9y = 5$

$12.6 + y = 9y = 3$

$13.y + 9 = 10y = 1$

$14.y + 6 = 12y = 6$

$15.3 + y = 9y = 6$

$16.y + 5 = 6y = 1$

$17.y + 1 = 3y = 2$

$18.5 + y = 14y = 9$

$19.9 + y = 12y = 3$

$20.1 + y = 4y = 3$

$21.3 + y = 7y = 4$

$22.y + 9 = 16y = 7$

$23.3 + y = 6y = 3$

$24.y + 1 = 7y = 6$

$25.7 + y = 8y = 1$

$26.5 + y = 11y = 6$

$27.y + 7 = 13y = 6$

$28.y + 2 = 8y = 6$

$29.y + 1 = 9y = 8$

$30.4 + y = 11y = 7$

$31.8 + y = 9y = 1$

$32.y + 2 = 11y = 9$

$33.1 + y = 6y = 5$

$34.y + 5 = 9y = 4$

$35.2 + y = 6y = 4$

$36.8 + y = 14y = 6$

$37.y + 3 = 10y = 7$

$38.5 + y = 8y = 3$

$39.y + 8 = 15y = 7$

$40.9 + y = 13y = 4$

$41.2 + y = 10y = 8$

$42.y + 1 = 5y = 4$

$43.y + 4 = 12y = 8$

$44.y + 3 = 8y = 5$

$45.y + 7 = 15y = 8$

$46.1 + y = 8y = 7$

$47.y + 8 = 12y = 4$

$48.9 + y = 18y = 9$

$49.6 + y = 11y = 5$

$50.y + 5 = 7y = 2$

$51.y + 2 = 4y = 2$

$52.y + 5 = 10y = 5$

$53.y + 8 = 17y = 9$

$54.y + 6 = 14y = 8$

$55.4 + y = 10y = 6$

$56.7 + y = 12y = 5$

$57.y + 2 = 5y = 3$

$58.y + 6 = 7y = 1$

$59.y + 7 = 16y = 9$

$60.7 + y = 14y = 7$

$61.6 + y = 15y = 9$

$62.y + 5 = 13y = 8$

$63.1 + y = 2y = 1$

$64.y + 4 = 7y = 3$

$65.y + 3 = 11y = 8$

$66.y + 8 = 13y = 5$

$67.y + 2 = 7y = 5$

$68.7 + y = 10y = 3$

$69.y + 8 = 16y = 8$

$70.6 + y = 10y = 4$

$71.y + 3 = 4y = 1$

$72.y + 6 = 13y = 7$

# Flvs Pre Algebra Answer Key

**Mary P. Dolciani, Robert H.  
Sorgenfrey, John Alexander Graham**

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web the balanced chemical equation for this reaction is  $\text{Mg} + 2\text{HNO}_3 \text{aq} \rightarrow \text{Mg}(\text{NO}_3)_2 \text{aq} + \text{H}_2 \text{g}$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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 $\text{mg}$  s magnesium 2  $\text{hno}_3$  l nitric acid  $\text{mg}$  no 3 2 s magnesium nitrate h 2 g hydrogen gas oxygen and hydrogen get  
 automatically balanced by balancing nitrogen

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 g  $\text{h}_2\text{o}$  l b  $\text{mg}^{2+}$  aq  $\text{co}_3^{2-}$  aq 2h aq 2no3 aq  $\text{mg}$  write the balanced equation when nitric acid reacts with calcium hydroxide to  
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 you balance magnesium and nitric acid you ll need to count all of atoms on both directions of the chemical equation to  
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 chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement  
 reaction the  $\text{mg}$  replaces the h in the  $\text{hno}_3$

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 each compound to balance  $\text{mg}$   $\text{hno}_3$   $\text{mg}$  no3 2  $\text{no}_2$   $\text{h}_2\text{o}$  you ll need to be

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