

# $d^2sp^3$ Hybridization



- $d^2sp^3$  hybridization enables an octahedral arrangement for six pairs of electrons surrounding a central atom.
  - Forms from two  $d$  orbitals, one  $s$  orbital, and three  $p$  orbitals.
- Examples
  - $\text{SF}_6$
  - $\text{XeF}_4$  (Xe has two lone pairs)



# Evan P Silberstein 20answers Atomic Structure

**Wenbin Ji**



**Evan P Silberstein 20answers Atomic Structure:**

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