

Chemistry EOC Review – Answer Key

- 1) a. 68,200,000 cg b. 5000 mm c. 0.0125 kL d. 0.548 dm
- 2) ~21.5 mL
- 3) production of heat & light, formation of a gas, formation of a precipitate, color change
- 4) $33.64 \text{ g} / 4.28 \text{ ml} = 7.86 \text{ g/ml} = \text{Fe}$
- 5) m.p. $-183^{\circ}\text{C} = \text{CH}_4$ (methane)
- 6) H_2 – splint test (burning splint ignites hydrogen)
 O_2 – splint test (oxygen reignites glowing splint)
 H_2O – cobalt chloride paper (paper changes from blue to pink in the presence of water vapor)
 CO_2 – splint test (carbon dioxide extinguishes burning splint);
limewater test (limewater turns cloudy when carbon dioxide is bubbled through it,
producing CaCO_3)
- 7) The protons are positively charged particles found inside the nucleus along with the neutrons, which have no charge. The total number of these nucleons in an atom is the mass # / atomic mass of the atom. The electrons are the negatively charged particles found orbiting in "clouds" around the nucleus. Electrons can be gained by the atom to form anions or lost to form cations.
- 8) The thing that makes any element on the periodic table different from all other elements is the number of protons in its nucleus. Isotopes are atoms of the same type of element which have different masses because they have different numbers of neutrons.
- 9) a. $^{39}_{19}\text{K}^+$
b. $^{57}_{26}\text{Fe}$
c. $^{33}_{16}\text{S}^{2-}$
- 10)

| <u>chemical symbol</u> | <u>atomic number</u> | <u>mass number</u> | <u># of protons</u> | <u># of electrons</u> | <u># of neutrons</u> | <u>net charge</u> |
|------------------------|----------------------|--------------------|---------------------|-----------------------|----------------------|-------------------|
| K^+ | 19 | 42 | 19 | 18 | 23 | +1 |
| Sr | 38 | 88 | 38 | 38 | 50 | 0 |
| F^- | 9 | 20 | 9 | 10 | 11 | -1 |
| P | 15 | 31 | 15 | 15 | 16 | 0 |
| Ca^{2+} | 20 | 42 | 20 | 18 | 22 | +2 |

- 11) uranium – 238 $238 - 92 = 146$ neutrons
- 12) (1) 199.59g
(2) They contain the same # of protons (they are the same element), but different #s of neutrons (electrons don't matter, b/c they have virtually no mass), & thus different masses.

Eoc Review Quiz Answer Key Chemistry

Arshad Iqbal



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covalent crystals properties of crystalline solids properties of positive rays quantum numbers quantum theory relative abundance Rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der Waals equation vapor pressure and spectrum

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Polymerization MCQ Chapter 24 Rates of Reaction MCQ Chapter 25 Reaction Kinetics MCQ Chapter 26 Redox Reactions and Electrolysis MCQ Chapter 27 States of Matter MCQ Chapter 28 Transition Elements MCQ The Alcohols and Esters MCQ PDF e Book Chapter 1 practice test to solve MCQ questions on Introduction to alcohols and alcohols reactions The Atomic Structure and Theory MCQ PDF e Book Chapter 2 practice test to solve MCQ questions on Atom facts elements and atoms number of nucleons protons electrons and neutrons The Benzene Chemical Compound MCQ PDF e Book Chapter 3 practice test to solve MCQ questions on Introduction to benzene arenes reaction phenol and properties and reactions of phenol The Carbonyl Compounds MCQ PDF e Book Chapter 4 practice test to solve MCQ questions on Introduction to carbonyl compounds aldehydes and ketone testing nucleophilic addition with HCN preparation of aldehydes and ketone reduction of aldehydes and ketone The Carboxylic Acids and Acyl Compounds MCQ PDF e Book Chapter 5 practice test to solve MCQ questions on Acidity of carboxylic acids acyl chlorides ethanoic acid and reactions to form tri iodomethane The Chemical Bonding MCQ PDF e Book Chapter 6 practice test to solve MCQ questions on Chemical bonding types chemical bonding electron pair bond angle bond energy bond energy bond length bonding and physical properties bonding energy repulsion theory covalent bonding covalent bonds double covalent bonds triple covalent bonds electron pair repulsion and bond angles electron pair repulsion theory enthalpy change of vaporization intermolecular forces ionic bonding ionic bonds and covalent bonds ionic bonds metallic bonding metallic bonding and delocalized electrons number of electrons sigma bonds and pi bonds sigma bonds pi bonds s orbital and p orbital Van der Waals forces and contact points The Chemistry of Life MCQ PDF e Book Chapter 7 practice test to solve MCQ questions on Introduction to chemistry enzyme specificity enzymes reintroducing amino acids and proteins The Electrode Potential MCQ PDF e Book Chapter 8 practice test to solve MCQ questions on Electrode potential cells and batteries E Plimsoll values electrolysis process measuring standard electrode potential quantitative electrolysis redox and oxidation The Electrons in Atoms MCQ PDF e Book Chapter 9 practice test to solve MCQ questions on Electronic configurations electronic structure evidence ionization energy periodic table simple electronic structure sub shells and atomic orbitals The Enthalpy Change MCQ PDF e Book Chapter 10 practice test to solve MCQ questions on Standard enthalpy changes bond energies enthalpies Hess law introduction to energy changes measuring enthalpy changes The Equilibrium MCQ PDF e Book Chapter 11 practice test to solve MCQ questions on Equilibrium constant expression equilibrium position acid base equilibria chemical industry equilibria ethanoic acid gas reactions equilibria and reversible reactions The Group IV MCQ PDF e Book Chapter 12 practice test to solve MCQ questions on Introduction to group IV metallic character of group IV elements ceramic silicon oxide covalent bonds properties variation in group IV relative stability of oxidation states and tetra chlorides The Groups II and VII MCQ PDF e Book Chapter 13 practice test to solve MCQ questions on Atomic number of group II metals covalent bonds density of group II elements disproportionation fluorine group II elements and reactions group VII elements and reactions halogens and compounds ionic

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Multiple Choice Questions 39 MCQs Energy from Chemicals Multiple Choice Questions 41 MCQs Experimental Chemistry Multiple Choice Questions 18 MCQs Methods of Purification Multiple Choice Questions 84 MCQs Particles of Matter Multiple Choice Questions 45 MCQs Redox Reactions Multiple Choice Questions 42 MCQs Salts and Identification of Ions and Gases Multiple Choice Questions 61 MCQs Speed of Reaction Multiple Choice Questions 35 MCQs Structure of Atom Multiple Choice Questions 52 MCQs The chapter Acids and Bases MCQs covers topics of acid rain acidity needs water acidity or alkalinity acids properties and reactions amphoteric oxides basic acidic neutral and amphoteric chemical formulas chemical reactions chemistry reactions college chemistry mineral acids general properties neutralization ordinary level chemistry organic acid pH scale acid and alkali properties bases and reactions strong and weak acids and universal indicators The chapter Chemical Bonding and Structure MCQs covers topics of ions and ionic bonds molecules and covalent bonds evaporation ionic and covalent substances ionic compounds crystal lattices molecules and macromolecules organic solvents polarization and transfer of electrons The chapter Chemical Formulae and Equations MCQs covers topics of chemical formulas chemical equations atomic mass ionic equations chemical reactions chemical symbols mixtures and compounds molar mass percent composition of elements reactants relative molecular mass valency and chemical formula and valency table The chapter Electricity MCQs covers topics of chemical to electrical energy applications of electrolysis reactions conductors and non conductors dry cells electrical devices circuit symbols electrolytes non electrolytes organic solvents polarization and valence electrons The chapter Electricity and Chemicals MCQs covers topics of chemical to electrical energy dry cells electrolyte non electrolyte and polarization The chapter Elements Compounds and Mixtures MCQs covers topics of elements compounds mixtures molecules atoms and symbols for elements

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Table of Contents Eoc Review Quiz Answer Key Chemistry

1. Understanding the eBook Eoc Review Quiz Answer Key Chemistry
 - The Rise of Digital Reading Eoc Review Quiz Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Eoc Review Quiz Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Eoc Review Quiz Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Eoc Review Quiz Answer Key Chemistry

- Personalized Recommendations
- Eoc Review Quiz Answer Key Chemistry User Reviews and Ratings
- Eoc Review Quiz Answer Key Chemistry and Bestseller Lists
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