

CLASS SET

Mixed Stoichiometry Problems

1. How many moles of H_2 would be required to completely react with O_2 to produce 5 moles of water? **5 mol H_2**
2. $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
 - a. Balance this equation
 - b. What mass of H_2SO_4 would be required to react with 0.75 mol of NaOH ? **37g**
3. What mass of NO_2 is formed when NO reacts with 384 g of O_2 ? **1104g**
4. If 30 g of H_2SO_4 reacts with aluminum hydroxide in a double replacement reaction, what mass of water is produced? **11.0g**
5. If 10 L of CH_4 combusts with O_2 , what volume of water will be produced? **20 L H_2O**
6. Tin(II) fluoride, SnF_2 , is used in some toothpastes. It is made by the reaction of tin with hydrogen fluoride according to the following equation.
$$\text{Sn} + 2\text{HF} \rightarrow \text{SnF}_2 + \text{H}_2$$

How many grams of SnF_2 are produced from the reaction of 30 g of HF with Sn ? **118 g**
7. In a spacecraft, the carbon dioxide exhaled by astronauts can be removed by its reaction with lithium hydroxide, LiOH , according to the following chemical equation.
$$\text{CO}_2(\text{g}) + 2\text{LiOH}(\text{s}) \rightarrow \text{Li}_2\text{CO}_3(\text{s}) + \text{H}_2\text{O}(\text{l})$$

If 448 L of CO_2 is exhaled, the average amount exhaled by a person each day, how much (in grams) $\text{Li}_2\text{CO}_3(\text{s})$ is produced? **1480g Li_2CO_3**
8. When 9.8g of aluminum oxide decomposes, how many grams of aluminum metal are produced? **5.2 g Al**
9. How many grams of iodine are produced when 0.72 mol of fluorine react with potassium iodide?
10. How many grams of sodium are required to react with water to produce 5.0g of sodium hydroxide? ($\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$)
11. How many grams of H_2 are produced in the reaction in #10?
12. What volume would be occupied by the resulting hydrogen at STP?
13. 20.0g of K react with H_2O to produce KOH and H_2 gas. How many liters would the hydrogen gas occupy at STP?
14. If 30.2g of Al react with HCl to produce AlCl_3 and H_2 gas, how many liters of H_2 are produced at STP?
15. How many liters of HCl are produced by the reaction of 5.7L of H_2 with Cl_2 ?

Mixed Stoichiometry Review Worksheet Answers

Julie C. Gilbert



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Table of Contents Mixed Stoichiometry Review Worksheet Answers

1. Understanding the eBook Mixed Stoichiometry Review Worksheet Answers
 - The Rise of Digital Reading Mixed Stoichiometry Review Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Stoichiometry Review Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Stoichiometry Review Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Stoichiometry Review Worksheet Answers
 - Personalized Recommendations
 - Mixed Stoichiometry Review Worksheet Answers User Reviews and Ratings

- Mixed Stoichiometry Review Worksheet Answers and Bestseller Lists
- 5. Accessing Mixed Stoichiometry Review Worksheet Answers Free and Paid eBooks
 - Mixed Stoichiometry Review Worksheet Answers Public Domain eBooks
 - Mixed Stoichiometry Review Worksheet Answers eBook Subscription Services
 - Mixed Stoichiometry Review Worksheet Answers Budget-Friendly Options
- 6. Navigating Mixed Stoichiometry Review Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mixed Stoichiometry Review Worksheet Answers Compatibility with Devices
 - Mixed Stoichiometry Review Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Stoichiometry Review Worksheet Answers
 - Highlighting and Note-Taking Mixed Stoichiometry Review Worksheet Answers
 - Interactive Elements Mixed Stoichiometry Review Worksheet Answers
- 8. Staying Engaged with Mixed Stoichiometry Review Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Stoichiometry Review Worksheet Answers
- 9. Balancing eBooks and Physical Books Mixed Stoichiometry Review Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Stoichiometry Review Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mixed Stoichiometry Review Worksheet Answers
 - Setting Reading Goals Mixed Stoichiometry Review Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mixed Stoichiometry Review Worksheet Answers
 - Fact-Checking eBook Content of Mixed Stoichiometry Review Worksheet Answers
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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