



10. Which of the following are endothermic processes?

- A. ice melting
- B. water condensing
- C. freezing
- D. sublimation

11. Heat absorbed per mole at 25°C, positive temperature, is:

12. Which of the following is a statement of the first law of thermodynamics?

- A. the negative change in enthalpy corresponds to an exothermic process
- B. energy lost by the system must be gained by the surroundings, neither of which are created
- C. temperature is a measure of average kinetic energy
- D. enthalpy is a state function

13. The specific enthalpy of combustion of glucose, C₆H₁₂O₆, is negative. This means that the product is _____.

- A. a hydrocarbon with the same number of atoms as glucose
- B. a hydrocarbon with the same number of atoms as glucose and oxygen
- C. a hydrocarbon with the same number of atoms as glucose
- D. a hydrocarbon with the same number of atoms as glucose and oxygen
- E. none of these is a true statement

14. The reaction:



is _____ and therefore heat is _____ during reaction.

- A. endothermic, released
- B. exothermic, absorbed
- C. endothermic, released
- D. exothermic, absorbed
- E. thermodynamic value released and absorbed

15. Heat evolved at 25°C are observed when 0.5 g of a solid is completely oxidized in the reaction:
 2C₂H₅OH + 3O₂ → 2CO₂ + 3H₂O ΔH = -1366 kJ



- A. 1366 kJ
- B. 683 kJ
- C. 683 or 1366 kJ
- D. 1366 kJ
- E. 683 kJ



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Robert R. Redfield



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Table of Contents Mcps Honors Chemistry Exam Guide

1. Understanding the eBook Mcps Honors Chemistry Exam Guide
 - The Rise of Digital Reading Mcps Honors Chemistry Exam Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcps Honors Chemistry Exam Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcps Honors Chemistry Exam Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcps Honors Chemistry Exam Guide
 - Personalized Recommendations
 - Mcps Honors Chemistry Exam Guide User Reviews and Ratings
 - Mcps Honors Chemistry Exam Guide and Bestseller Lists

5. Accessing Mcps Honors Chemistry Exam Guide Free and Paid eBooks
 - Mcps Honors Chemistry Exam Guide Public Domain eBooks
 - Mcps Honors Chemistry Exam Guide eBook Subscription Services
 - Mcps Honors Chemistry Exam Guide Budget-Friendly Options
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 - Mcps Honors Chemistry Exam Guide Compatibility with Devices
 - Mcps Honors Chemistry Exam Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcps Honors Chemistry Exam Guide
 - Highlighting and Note-Taking Mcps Honors Chemistry Exam Guide
 - Interactive Elements Mcps Honors Chemistry Exam Guide
8. Staying Engaged with Mcps Honors Chemistry Exam Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcps Honors Chemistry Exam Guide
9. Balancing eBooks and Physical Books Mcps Honors Chemistry Exam Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcps Honors Chemistry Exam Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcps Honors Chemistry Exam Guide
 - Setting Reading Goals Mcps Honors Chemistry Exam Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcps Honors Chemistry Exam Guide
 - Fact-Checking eBook Content of Mcps Honors Chemistry Exam Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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