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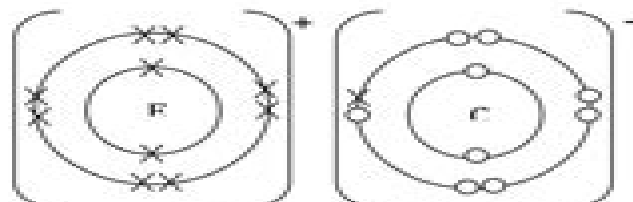
CHEMISTRY PAPER 2 (233/2)

FORM FOUR (4)

TIME: 2 HOURS

MARKING SCHEME

1. 1 (a) Alkaline-earth metals (1mk)  
 (b) B (1mk)  
 (c) (i)  $A_2B_3$  (½mk) (ii)  $H_2SO_4$  (½mk)  
 (d)



Distribution of electrons (1mk) and charges (1mk)

No label (0mk)

- (e) D has a smaller atomic size than C // C has larger atomic size than D (1mk)  
 D has more protons than C // D has stronger nuclear charge than C (1mk)  
 (f) H (½mk). Loses electrons most readily / largest atomic radius / weakest nuclear charge (explanation ½mk)  
 (g)  $2A_{(s)} + 6HCl_{(aq)} \longrightarrow 2ACl_{3(aq)} + 3H_{2(g)}$  1mk  
 Moles of A =  $\frac{0.081}{27} = 0.003$  moles ½mk  
 Mole ratio 1:3  
 Moles of HCl =  $0.003 \times 3 = 0.009$  moles ½mk  
 $\frac{0.009 \times 1000}{20}$  ½mk = 0.45M ½mk  
 (h) The melting point increase with increase in atomic number. 1mk

This is due to increase in the strength of van der waal's forces (inter-molecular forces) 1mk

# Marking Scheme Chemistry June 2013 Paper 2

**Iseb Iseb**



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