



EMC® Open Migrator/LM for Windows

Version 3.12

PRODUCT GUIDE

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AW Chickering



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 time the total pressure $p_a = x_a p_{\text{total}}$ mole fraction $x_a = \frac{\text{moles of substance a}}{\text{total moles}}$
 a is often denoted by x_a and

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 present is equal to one $x_a + x_b = 1$ when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is
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 ways to represent the concentration of the solution it is defined as the number of molecules of component a divided by the
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 mass $C_6H_6 = 78 \text{ g/mol}$ hence C_6H_6 's number of moles $\frac{30}{78} = 0.3846 \text{ mol}$ molar mass of carbon tet $CCl_4 = 154 \text{ g/mol}$

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 g of water calculate the molarity and molality of the solution given $N = 14, H = 1, C = 12, O = 16$ given mass of solute urea 11.11 g mass
 of solvent water 100 g = 0.1 kg to find molarity of solution solution

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 of $CH_3OH = \frac{0.173}{0.173 + 2.23} = 0.073$ to solve more examples on mole fraction formula and practice more
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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the O_2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature

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