

### 1.1<sup>a</sup>. Concentration of a gas mixture.

A mixture of noble gases [helium (1), argon (2), krypton (3), and xenon (4)] is at a total pressure of 200 kPa and a temperature of 400 K. If the mixture has equal mole fractions of each of the gases, determine:

a) The composition of the mixture in terms of mass fractions.

Solution:

Basis: 100 kmole of the mixture

$$\begin{aligned}
 \text{ORIGIN} &:= 1 & \text{kmole} &:= 1000 \cdot \text{mole} \\
 \text{moles} &:= \begin{pmatrix} 25 \\ 25 \\ 25 \\ 25 \end{pmatrix} \cdot \text{kmole} & \text{MW} &:= \begin{pmatrix} 4 \\ 40 \\ 83.8 \\ 131.3 \end{pmatrix} \frac{\text{kg}}{\text{kmole}} & \text{mass} &:= \overrightarrow{(\text{moles} \cdot \text{MW})} \\
 \text{mass} &:= \begin{pmatrix} 100 \\ 1 \times 10^3 \\ 2.095 \times 10^3 \\ 3.283 \times 10^3 \end{pmatrix} \text{kg} & \text{total} &:= \sum_{i=1}^4 \text{mass}_i & \text{total} &:= 6.478 \times 10^3 \text{ kg} \\
 \text{mass\_fraction} &:= \frac{\text{mass}}{\text{total}} & \text{mass\_fraction} &:= \begin{pmatrix} 0.015 \\ 0.154 \\ 0.323 \\ 0.507 \end{pmatrix}
 \end{aligned}$$

b) The average molecular weight of the mixture.

Solution

$$\begin{aligned}
 \text{kmole} &:= 1000 \cdot \text{mole} \\
 M_{av} &:= \frac{\text{total}}{100 \cdot \text{kmole}} & M_{av} &= 0.065 \frac{\text{kg}}{\text{mol}}
 \end{aligned}$$

c) The total molar concentration.

Solution

$$\begin{aligned}
 P &:= 200 \cdot \text{kPa} & T &:= 400 \cdot \text{K} & \text{kPa} &:= 1000 \cdot \text{Pa} \\
 c &:= \frac{P}{R \cdot T} & R &:= 8.314 \cdot \text{kPa} \cdot \text{m}^3 \cdot \text{kmole}^{-1} \cdot \text{K}^{-1} \\
 c &= 60.14 \frac{\text{mol}}{\text{m}^3}
 \end{aligned}$$

d) The mass density.

Solution

$$\begin{aligned}
 \rho &:= c \cdot M_{av} & \rho &= 3.896 \frac{\text{kg}}{\text{m}^3}
 \end{aligned}$$

# Manual Solution Of Treybal

**M Mark**



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