

1. Determine the kinds of intermolecular forces (dispersion, dipole-dipole, H-bonding) that are present in each of the following elements or compounds.

- a. Kr dispersion
- b. NCl_3 dispersion, dipole-dipole
- c. SiH_4 dispersion
- d. HF dispersion, dipole-dipole, H-bonding
- e. N_2 dispersion
- f. NH_3 dispersion, dipole-dipole, H-bonding
- g. CO dispersion, dipole-dipole
- h. CO_2 dispersion,

2. Arrange the following compounds in order of increasing boiling point, explain your reasoning.



3. Which one of the molecules would require the most energy to break the IMFs? How would you rank them in term of their boiling points?






4. Which one of the following compounds is more likely to be a liquid at room temperature, and why?



While all three compounds have similar molar mass, and all have dispersion and dipole-dipole forces, only hydrogen peroxide has H-bonding, thus likely to have the highest b.p. amongst the three compounds.

5. The surface tension and viscosity values for diethyl ether, acetone, ethanol and ethylene glycol are shown here:

Compound	Molecule	Surface Tension (mN/m)	Viscosity (mPa s)
diethyl ether $\text{C}_2\text{H}_5\text{OC}_2\text{H}_5$		17	0.22
ethanol $\text{C}_2\text{H}_5\text{OH}$		22	1.07
ethylene glycol $\text{CH}_2(\text{OH})\text{CH}_2(\text{OH})$		48	16.1

- a. Explain the differences in viscosity in terms of the size and shape of the molecules and their IMFs.
The molecules with the most IMFs will have the highest viscosity. Ethylene glycol is the largest molecule and has two H-bonding sites, whereas ethanol is small and has one H-bonding. Diethylether is the least viscous since it does not have H-bonding.
- b. Explain their differences in surface tension in terms of the size and shape of the molecules and their IMFs.
Similar trend is observed for surface tension, where the molecule with the most IMFs (ethylene glycol) will exhibit the highest surface tension.

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Table of Contents Intermolecular Forces And Organics Worksheet Answer Key

1. Understanding the eBook Intermolecular Forces And Organics Worksheet Answer Key
 - The Rise of Digital Reading Intermolecular Forces And Organics Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Intermolecular Forces And Organics Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Intermolecular Forces And Organics Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Intermolecular Forces And Organics Worksheet Answer Key

- Personalized Recommendations
 - Intermolecular Forces And Organics Worksheet Answer Key User Reviews and Ratings
 - Intermolecular Forces And Organics Worksheet Answer Key and Bestseller Lists
5. Accessing Intermolecular Forces And Organics Worksheet Answer Key Free and Paid eBooks
 - Intermolecular Forces And Organics Worksheet Answer Key Public Domain eBooks
 - Intermolecular Forces And Organics Worksheet Answer Key eBook Subscription Services
 - Intermolecular Forces And Organics Worksheet Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Intermolecular Forces And Organics Worksheet Answer Key Compatibility with Devices
 - Intermolecular Forces And Organics Worksheet Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Intermolecular Forces And Organics Worksheet Answer Key
 - Highlighting and Note-Taking Intermolecular Forces And Organics Worksheet Answer Key
 - Interactive Elements Intermolecular Forces And Organics Worksheet Answer Key
 8. Staying Engaged with Intermolecular Forces And Organics Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Intermolecular Forces And Organics Worksheet Answer Key
 9. Balancing eBooks and Physical Books Intermolecular Forces And Organics Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Intermolecular Forces And Organics Worksheet Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Intermolecular Forces And Organics Worksheet Answer Key
 - Setting Reading Goals Intermolecular Forces And Organics Worksheet Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Intermolecular Forces And Organics Worksheet Answer Key

- Fact-Checking eBook Content of Intermolecular Forces And Organics Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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