Worksheet #8 Intermolecular Forces

Chem 102

 Determine the kinds of intermolecular forces (dispersion, dipole-dipole, H-bonding) that are present in each of the following elements or compounds.

40.00	Committee of the committee of the	distribution of contributions.
A.,	Kr	dispersion
b.,	NCI	dispersion, dipole-dipole
C.	SIH ₄	dispersion
d.	HF	dispersion, dipole-dipole, H-bonding
0.	Nz	dispersion
f.	NHa	dispersion, dipole-dipole, H-bonding
g.	CO	dispersion, dipole-dipole
Sec. 1	CO ₄	dispersion.

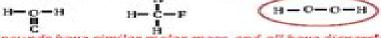
2. Arrange the following compounds in order of increasing boiling point, explain your reasoning.

CH4	CH ₂ CH ₂ CI	CH ₃ CH ₃	CH ₂ CH ₂ OH
CH ₄ dispersion	< CHbCHs dispersion (but heavi- than CH ₄)	dispersion ler dipole-dip	+ dispersion +

3. Which one of the molecules would require the most energy to break the IMFs? How would you rank them in term of their boiling point?



4. Which one of the following compounds is more likely to be a liquid at room temperature, and why?



While all three compounds have similar motor mass, and all have dispersion and dipole-dipole forces, only hydrogen peroxide has H-bonding, thus likely to have the highest b.p. amongst the three compounds.

5. The surface tension and viscosity values for diethyl ether, acetone, ethanol and ethylene glycol are shown here:

Compound	Michecule	Surface Tension (mhim)	(mPa n)	
diethyl ether C ₃ H ₆ OC ₃ H ₆	and the same	3.2	0.22	
ethanol C ₂ H ₃ OH	48	22	1.07	
ethylene glycol CH ₂ (OH)CH ₂ (OH)	34	48	16.1	

Explain the differences in viscosity in terms of the size and shape of the molecules and their IMFs.

The melecules with the most IMFs will have the highest viscosity. Ethylene glycell is the largest molecule and has two H-bending sites, whereas ethanel is small and has one -H-bending. Diethylether is the least viscous since it does not have H-bending.

Explain their differences in surface tension in terms of the size and shape of the molecules and their IMFs.

Similar trend is observed for surface tension, where the molecule with the most IMFS (ethylene glycol) will exhibit the highest surface tension.

<u>Intermolecular Forces And Organics Worksheet Answer</u> <u>Key</u>

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