

⇒ Heisenberg uncertainty principle :

$$\Delta x \cdot \Delta p \geq \frac{h}{4\pi}$$

⇒ Number of orbitals in a shell = n^2

⇒ Number of electrons in a shell = $2n^2$

⇒ Number of radial nodes = $(n - l - 1)$

Chemical Bonding

$$H = \frac{1}{2}[V + X - C + A]$$

(H = number of Hybrid orbitals)

$$\Rightarrow \text{Bond order} = \frac{1}{2}[N_b - N_a]$$

States of matter

$$\Rightarrow \text{Boyle's law : } P_1 V_1 = P_2 V_2$$

$$\Rightarrow \text{Charles law : } \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\Rightarrow \text{Ideal gas equation : } PV = nRT$$

$$\text{Dalton's law : } \Rightarrow P_{\text{total}} = P_1 + P_2 + P_3 + \dots$$

$$\Rightarrow d_{\text{gas}} = \frac{PM}{RT}$$

$$\text{Graham's Law} \Rightarrow \frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$$

$$\Rightarrow \text{Vander waal's equation : } \left(P + \frac{an^2}{V^2} \right) (V - nb) = nRT$$

$$\Rightarrow C_p - C_v = nRT$$

$$\Rightarrow \text{Kinetic gas equation : } PV = \frac{1}{3} Mu^2$$

Chemical Thermodynamics

$$\Rightarrow \Delta U = q_p, \Delta H = q_p$$

$$\Rightarrow \Delta H = \Delta U + \Delta n_g RT$$

$$\Rightarrow C = \frac{q}{\Delta T} \quad (C = \text{Heat capacity})$$

$$\Rightarrow \Delta U = q + w \quad (\text{First law})$$

$$\Rightarrow \Delta H = \sum \Delta_f H^\circ_{\text{products}} - \sum \Delta_f H^\circ_{\text{reactants}}$$

$$\Rightarrow \Delta G = \Delta H - T\Delta S$$

$$\Rightarrow \Delta G^\circ = -RT \ln K$$

Equilibrium

$$\Rightarrow K_{eq} = \frac{K_c}{K_p}$$

$$\Rightarrow K_p = K_c (RT)^{\Delta n}$$

$$\Rightarrow K_w = [H^+][OH^-]$$

$$\Rightarrow pH = -\log [H^+]$$

$$\Rightarrow pK_w = pH + pOH$$

$$\Rightarrow Ka = C\alpha^2$$

$$\Rightarrow K_w = K_a \times K_b$$

$$\Rightarrow K_w = [A^-]_e [B^+]_e$$

Solid State

$$\Rightarrow \text{Density} = \frac{Z \times M}{a^3 \times N_A} \text{ g/cm}^3$$

$$\Rightarrow \text{Packing fraction : Simple cubic} = \frac{\pi}{6}, \text{ FCC} = \frac{\sqrt{2}\pi}{6}, \text{ BCC} = \frac{\sqrt{3}\pi}{8}$$

$$\Rightarrow \text{Bragg's equation } d = n\lambda \sin \theta$$

Solution

$$\Rightarrow \text{Henry's law } X_{\text{gas}} = K_H P$$

$$\Rightarrow \text{Raoult's law : v.p. of solution} = x_A P_A^\circ + x_B P_B^\circ$$

$$\Rightarrow \text{Relative lowering of V.P. : } \frac{P_A^\circ - P_A}{P_A^\circ} = X_{\text{solvent}}$$

$$\Rightarrow \Delta T_b = i K_b m$$

$$\Rightarrow \Delta T_f = i K_f m$$

$$\Rightarrow \pi = CRT$$

$$\Rightarrow i = \frac{\text{no. of particle of solute in solution}}{\text{no. of particles of pure solute}}$$

Electrochemistry

$$\Rightarrow E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$$

$$\Rightarrow \Delta G^\circ = -nF E^\circ_{\text{cell}}$$

Hsc Chemistry Formula Sheet

Brendan G. Carr

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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