

Name: _____ Date: _____ Lab Section: _____

Experiment 2: Molecular Models REPORT SHEET

Pre-lab Study Questions:

- Objectives of the experiment (in your own words):
- What kinds of electrons form the covalent bonds in molecules?
- Is a polar molecule symmetrical or asymmetrical?

FORMULA OF MOLECULE	LEWIS STRUCTURE	3-D SHAPE w/ lone pairs	MOLECULAR GEOMETRY	MOLECULAR POLARITY
example: CH ₄	$ \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array} $		tetrahedral	nonpolar
H ₂ O	$ \begin{array}{c} \text{H} \quad \text{H} \\ \diagdown \quad \diagup \\ \text{O} \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array} $	$ \begin{array}{c} \text{H} \quad \text{H} \\ \diagdown \quad \diagup \\ \text{O} \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array} $	tetrahedral	polar
H ₂ O ⁺	$ \begin{array}{c} \text{H} \quad \text{H} \\ \diagdown \quad \diagup \\ \text{O} \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array} $	$ \begin{array}{c} \text{H} \quad \text{H} \\ \diagdown \quad \diagup \\ \text{O} \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array} $	tetrahedral	nonpolar
C ₂ H ₆				
C ₂ H ₄	$ \begin{array}{c} \text{H} \quad \text{H} \\ \diagdown \quad \diagup \\ \text{C}=\text{C} \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array} $		trigonal planar	non polar
C ₂ H ₂				
CO ₂	$ \begin{array}{c} \text{O}=\text{C}=\text{O} \end{array} $		Linear	Non polar
H ₂ O ₂				
CH ₃ Cl				

Molecular Modeling Lab Report

LL Leslie

A decorative red circular graphic with a gradient, appearing as a semi-circle or a stylized 'C' shape, located to the right of the author's name.

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