

d - block: [Noble gas] $(n-1)d^{1-10}ns^{1-2}$

f - block: [Noble gas] $(n-2)f^{1-14}(n-1)d^{0-1}ns^2$

5. The ingredients are classified into three categories:

(i). Metals account for more than 78% of all known elements. Metals are the s - block, d - block, and f - block elements. Metals make up the higher components of the p - block.

(ii). The number of non-metals is less than twenty. (Non-metals include C , N , P , O , S , Se , H , F , Cl , Br , I , He , Ne , Ar , Kr , Xe , and Rn.)

(iii). Semimetals or metalloids are elements that exist on the boundary between metals and non-metals. Metalloids include B , Si , Ge , As , Se , Te , Po , and At .

6. The IUPAC has devised a new mechanism for giving newly discovered elements a temporary name. The name is derived directly from the element's atomic number. The names of the elements from atomic numbers 104 to 110 have been accepted by IUPAC.

7. Periodicity is defined as the occurrence of identical properties of the elements at regular intervals when the elements are grouped in the periodic table in order of increasing atomic numbers. The repetition of similar electronic configurations of the atom in the valence shell after certain specified periods is the cause of periodicity. These definite intervals are 2, 8, 8, 18, 18 and 32. These are referred to as "magic numbers." Periodicity can be seen in a variety of qualities that are either directly or indirectly related to the electronic configuration.

(i). Across a period, the effective nuclear charge increases.

(ii). Atomic radii tend to decrease as time passes.

(iii). As you progress through the groupings, the atomic radii tend to rise.

(iv). There are three types of atomic radius:

a. Covalent radius: The distance between the nuclei of two comparable atoms linked by a single covalent bond is half of this radius. Non-metals are typically described in this way.

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