

Exercise 14.90 with feedback

You may want to reference ([10 pages 535 - 543](#)) section 14.7 while completing this problem.

A 0.10 m aqueous solution of an ionic compound with the formula MX has a freezing point of $-2.3\text{ }^{\circ}\text{C}$.

Part A

Calculate the van't Hoff factor (i) for MX at this concentration.

Express your answer using two significant figures.



Calculator



Alpha Omega



Undo



Redo



Clear



Help

$i =$

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