

CHEMISTRY

CHEAT SHEET

CHEMICAL KINETICS

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DEFINITIONS:

- Rate of Reaction:** The change in concentration of a reactant or product per unit time.
- Rate of Disappearance:** The rate at which a reactant is consumed.
- Rate of Appearance:** The rate at which a product is formed.

Rate Laws:

- Rate Law:** An equation that relates the rate of a reaction to the concentration of the reactants.
- Rate Constant (k):** A proportionality constant that is unique to a particular reaction at a given temperature.
- Order of Reaction:** The sum of the exponents in the rate law.

Graphs:

- Rate vs. Time:** A graph showing the rate of reaction as a function of time. The rate decreases as time increases.
- Concentration vs. Time:** A graph showing the concentration of a reactant or product as a function of time. The concentration of a reactant decreases, while the concentration of a product increases.

Order	Rate Law	Half-life	Graph
Zero-order	$\text{Rate} = k$	$t_{1/2} = \frac{[A]_0}{k}$	
First-order	$\text{Rate} = k[A]$	$t_{1/2} = \frac{\ln 2}{k}$	
Second-order	$\text{Rate} = k[A]^2$	$t_{1/2} = \frac{1}{k[A]_0}$	

Integration of Rate Laws:

- Zero-order:** $[A]_0 - kt = [A]$
- First-order:** $\ln[A]_0 - kt = \ln[A]$
- Second-order:** $\frac{1}{[A]_0} - kt = \frac{1}{[A]}$

Arrhenius Equation:

$$k = Ae^{-E_a/RT}$$

Activation Energy (E_a): The minimum energy required for a reaction to occur.

Temperature: A measure of the average kinetic energy of the particles in a system.

Kinetics Study Guide

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