

# Intro to Acids and Bases

Name: \_\_\_\_\_  
 Class: \_\_\_\_\_  
 Date: \_\_\_\_\_

## Define each term: Definitions of acids and bases:

### Arrhenius Definitions:

- acid substances that release dissolved hydrogen ions ( $H^+$ ) (acids  $H^+$  gases formed in gases as the hydrogens form  $H_2$  and  $H_2O$  (acid  $H_2O$ ) and equivalent equivalents)
- base substances which release hydroxide ions ( $OH^-$ )

### Brønsted-Lowry Definitions:

- acid substances that donate a proton ( $H^+$ ) (acid donate)
- base substances which accept a proton ( $H^+$ ) (base receive)

Table 1: Representative reactions for  $H^+$  (acid) and  $OH^-$  (base) representative reactions:

Reaction	Eq.
1. $HCl(aq) + H_2O(l) \rightarrow H_3O^+(aq) + Cl^-(aq)$	$HCl \rightarrow H^+ + Cl^-$
2. $HNO_3(aq) + H_2O(l) \rightarrow H_3O^+(aq) + NO_3^-(aq)$	$HNO_3 \rightarrow H^+ + NO_3^-$
3. $H_2SO_4(aq) + H_2O(l) \rightarrow H_3O^+(aq) + HSO_4^-(aq)$	$H_2SO_4 \rightarrow H^+ + HSO_4^-$
4. $H_2SO_4(aq) + H_2O(l) \rightarrow H_3O^+(aq) + SO_4^{2-}(aq)$	$H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$
5. $NaOH(aq) + H_2O(l) \rightarrow Na^+(aq) + OH^-(aq)$	$NaOH \rightarrow Na^+ + OH^-$

## Critical Thinking Questions

- Are all the substances in Table 1 that are representative acids?  
 (a) no all some none
- Are all the substances in Table 1 that are representative bases?

### Class work:

- Are all the substances in Table 1 that are Brønsted-Lowry acids?

### Individual questions for students:

- Are all the substances in Table 1 that are Brønsted-Lowry bases?

### Group questions:

- Are possible reactions that are not acid-base reactions?

### Discussion questions and answers for class:

- Which substances are not Brønsted-Lowry acids/bases (different conditions)? Explain.

- Are all the substances that are Brønsted-Lowry acids/bases (different conditions)? Explain.

# Intro To Acids And Bases Chemquest 48

**Jean-Charles Joud, Marie-Geneviève  
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and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described      **Acids, Bases, and the Chemistry of the Covalent Bond** Calvin Anthony VanderWerf,1961      **How to Understand Acid-base** Peter A. Stewart,1981-01-01      Acid Base Indicators I. M. Kolthoff,2007-03-01 Many of the earliest books particularly those dating back to the 1900s and before are now extremely scarce and increasingly expensive We are republishing these classic works in affordable high quality modern editions using the original text and artwork      Physical Chemistry and Acid-base Properties of Surfaces Jean-Charles Joud,Marie-Geneviève Barthés-Labrousse,2015

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### **Table of Contents Intro To Acids And Bases Chemquest 48**

1. Understanding the eBook Intro To Acids And Bases Chemquest 48
  - The Rise of Digital Reading Intro To Acids And Bases Chemquest 48
  - Advantages of eBooks Over Traditional Books
2. Identifying Intro To Acids And Bases Chemquest 48
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Intro To Acids And Bases Chemquest 48
  - User-Friendly Interface
4. Exploring eBook Recommendations from Intro To Acids And Bases Chemquest 48
  - Personalized Recommendations
  - Intro To Acids And Bases Chemquest 48 User Reviews and Ratings
  - Intro To Acids And Bases Chemquest 48 and Bestseller Lists

5. Accessing Intro To Acids And Bases Chemquest 48 Free and Paid eBooks
  - Intro To Acids And Bases Chemquest 48 Public Domain eBooks
  - Intro To Acids And Bases Chemquest 48 eBook Subscription Services
  - Intro To Acids And Bases Chemquest 48 Budget-Friendly Options
6. Navigating Intro To Acids And Bases Chemquest 48 eBook Formats
  - ePub, PDF, MOBI, and More
  - Intro To Acids And Bases Chemquest 48 Compatibility with Devices
  - Intro To Acids And Bases Chemquest 48 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Intro To Acids And Bases Chemquest 48
  - Highlighting and Note-Taking Intro To Acids And Bases Chemquest 48
  - Interactive Elements Intro To Acids And Bases Chemquest 48
8. Staying Engaged with Intro To Acids And Bases Chemquest 48
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Intro To Acids And Bases Chemquest 48
9. Balancing eBooks and Physical Books Intro To Acids And Bases Chemquest 48
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Intro To Acids And Bases Chemquest 48
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Intro To Acids And Bases Chemquest 48
  - Setting Reading Goals Intro To Acids And Bases Chemquest 48
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Intro To Acids And Bases Chemquest 48
  - Fact-Checking eBook Content of Intro To Acids And Bases Chemquest 48
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
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