

## Practice Test Ch 3

## Stoichiometry

Name Key No.     

- Remember, this is practice - Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math" - common factors and rough estimation - do not do "long division" to try to get exact values. Remember it is a MC test, use the answers.
- Mark which questions you would like to "go over" when we get to school in September.

1. Balance the following equation:



The balanced equation shows that 1.00 mole of  $\text{NH}_3$  requires      mole(s) of  $\text{O}_2$ .

- a. 0.57  
b. 1.25  
c. 1.33  
**d. 1.75**  
e. 3.5

$$1 \text{ mol NH}_3 \left( \frac{7 \text{ mol O}_2}{4 \text{ mol NH}_3} \right) =$$

2. Write a balanced equation for the combustion of acetaldehyde,
- $\text{CH}_3\text{CHO}$
- .



When properly balanced, the equation indicates that      mole(s) of  $\text{O}_2$  are required for each mole of  $\text{CH}_3\text{CHO}$ .

- a. 1  
b. 2  
**c. 2.5**  
d. 3  
e. 6

$$1 \text{ mol CH}_3\text{CHO} \left( \frac{5 \text{ mol O}_2}{2 \text{ mol CH}_3\text{CHO}} \right) =$$

3. What is the total mass of products formed when 16 grams of
- $\text{CH}_4$
- is burned with excess oxygen?



$$16 \text{ g CH}_4 \left( \frac{1 \text{ mol CH}_4}{16 \text{ g CH}_4} \right) \left( \frac{2 \text{ mol O}_2}{1 \text{ mol CH}_4} \right) \left( \frac{32 \text{ g O}_2}{1 \text{ mol O}_2} \right) =$$

- a. 32 g  
b. 36 g  
c. 44 g  
d. 62 g  
**e. 80 g**

4. Write a balanced equation for the combustion of propane,
- $\text{C}_3\text{H}_8$
- .



When balanced, the equation indicates that      moles of  $\text{O}_2$  are required for each mole of  $\text{C}_3\text{H}_8$ .

- a. 1.5  
b. 3  
c. 2.5  
**d. 5**  
e. 8

5. Balance the following equation with the SMALLEST WHOLE NUMBER COEFFICIENTS possible. Select the number that is the sum of the coefficients in the balanced equation:



- a. 3  
b. 5  
c. 6  
**d. 7**  
e. 8

6. Calculate the mass of hydrogen formed when 27 g of aluminum reacts with excess hydrochloric acid according to the balanced equation below:



- a. 1.5 g  
b. 2.0 g  
**c. 3.0 g**  
d. 6.0 g  
e. 12 g

$$27 \text{ g Al} \left( \frac{1 \text{ mol Al}}{27 \text{ g Al}} \right) \left( \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \right) \left( \frac{2 \text{ g H}_2}{1 \text{ mol H}_2} \right) =$$

7. How many grams of nitric acid,
- $\text{HNO}_3$
- , can be prepared from the reaction of 138 g of
- $\text{NO}_2$
- with 54.0 g
- $\text{H}_2\text{O}$
- according to the equation below?



$$138 \text{ g NO}_2 \left( \frac{1 \text{ mol NO}_2}{46 \text{ g NO}_2} \right) \left( \frac{2 \text{ mol HNO}_3}{3 \text{ mol NO}_2} \right) \left( \frac{63 \text{ g HNO}_3}{1 \text{ mol HNO}_3} \right) = 126 \text{ g}$$

- a. 92  
b. 108  
**c. 126**  
d. 189  
e. 279

$$54 \text{ g H}_2\text{O} \left( \frac{1 \text{ mol H}_2\text{O}}{18 \text{ g H}_2\text{O}} \right) \left( \frac{2 \text{ mol HNO}_3}{1 \text{ mol H}_2\text{O}} \right) \left( \frac{63 \text{ g HNO}_3}{1 \text{ mol HNO}_3} \right) = 4 \times (63) \text{ g}$$

8. Which of the following statements is true?

- I. The molar mass of  $\text{CaCO}_3$  is 100.1 g  $\text{mol}^{-1}$ .  
II. 50 g of  $\text{CaCO}_3$  contains  $9 \times 10^{23}$  oxygen atoms.  
III. A 200 g sample of  $\text{CaCO}_3$  contains 2 moles of  $\text{CaCO}_3$ .

- a. I only  
b. II only  
c. III only  
**d. I and III only**  
e. I, II, and III

$$\frac{40}{12} = \frac{4.8}{100 \text{ g mol}^{-1}}$$

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