

1. Conceptual Questions

1.1. What is the difference between a **strong** and a **weak** acid?

Strong acids completely dissociate in water, while **weak** acids only partially dissociate. **Strong** acids have a **large** K_a value, while **weak** acids have a **small** K_a value.

1.2. What is the difference between a **strong** and a **weak** base? Give an example of each.

- a. **Strong** base: NaOH
- b. **Weak** base: NH_3
- c. **Strong** base: KOH
- d. **Weak** base: CH_3NH_2
- e. **Strong** base: LiOH
- f. **Weak** base: H_2O

Other examples of **strong** acids:

HCl , HNO_3 , H_2SO_4 , HClO_4

Other examples of **weak** acids:

CH_3COOH , H_2O , H_2S , H_2PO_4^-

1.3. What is a **buffer**?

A **buffer** is a solution that resists changes in pH when small amounts of **strong** acid or base are added.

A **buffer** is a solution that contains a **weak** acid and its **conjugate base** (or a **weak** base and its **conjugate acid**).

1.4. What is the **Henderson-Hasselbalch equation**? Give an example of a **buffer** solution.

The **Henderson-Hasselbalch equation** is:

$$\text{pH} = \text{pK}_a + \log \left(\frac{[\text{A}^-]}{[\text{HA}]}} \right)$$

Example: A **buffer** solution of **acetic acid** (CH_3COOH) and **acetate** (CH_3COO^-).

Example: A **buffer** solution of **ammonia** (NH_3) and **ammonium** (NH_4^+).

1.5. What is the **buffer capacity**?

1.6. What is the **buffer range**?

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